

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on  $K_a$  and  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                        |                             |                   |
|------------------------|-----------------------------|-------------------|
| 1. autoionization      | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization          | 5. $K_a$                    | 6. $K_b$          |
| 7. ionization constant | 8. $pH$                     |                   |

## Key Equation

- |                       |                       |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

## Part A Completion

Write the equation and/or the numerical value(s) of the constants indicated that are associated with each reaction. Be sure to include units where appropriate.

1. Water autoionization:  $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$   $K_w = 1.0 \times 10^{-14}$  at 25°C
2. Ionization of acetic acid:  $\text{CH}_3\text{COOH} \rightleftharpoons \text{CH}_3\text{COO}^- + \text{H}^+$   $K_a = 1.8 \times 10^{-5}$  at 25°C
3. Ionization of ammonia:  $\text{NH}_3 + \text{H}_2\text{O} \rightleftharpoons \text{NH}_4^+ + \text{OH}^-$   $K_b = 1.8 \times 10^{-5}$  at 25°C
4. Ionization of hydrofluoric acid:  $\text{HF} \rightleftharpoons \text{F}^- + \text{H}^+$   $K_a = 6.6 \times 10^{-4}$  at 25°C
5. Ionization of formic acid:  $\text{HCOOH} \rightleftharpoons \text{HCOO}^- + \text{H}^+$   $K_a = 1.8 \times 10^{-4}$  at 25°C
6. Ionization of nitrous acid:  $\text{HNO}_2 \rightleftharpoons \text{NO}_2^- + \text{H}^+$   $K_a = 4.5 \times 10^{-4}$  at 25°C
7. Ionization of carbonic acid:  $\text{H}_2\text{CO}_3 \rightleftharpoons \text{HCO}_3^- + \text{H}^+$   $K_a = 4.3 \times 10^{-7}$  at 25°C
8. Ionization of phosphoric acid:  $\text{H}_3\text{PO}_4 \rightleftharpoons \text{H}_2\text{PO}_4^- + \text{H}^+$   $K_a = 7.5 \times 10^{-3}$  at 25°C

9. Ionization of sulfuric acid:  $\text{H}_2\text{SO}_4 \rightleftharpoons \text{HSO}_4^- + \text{H}^+$   $K_a = 1.0 \times 10^{-2}$  at 25°C
10. Ionization of hydrochloric acid:  $\text{HCl} \rightleftharpoons \text{Cl}^- + \text{H}^+$   $K_a = 1.0 \times 10^7$  at 25°C
11. Ionization of hydrobromic acid:  $\text{HBr} \rightleftharpoons \text{Br}^- + \text{H}^+$   $K_a = 1.0 \times 10^7$  at 25°C
12. Ionization of hydroiodic acid:  $\text{HI} \rightleftharpoons \text{I}^- + \text{H}^+$   $K_a = 1.0 \times 10^7$  at 25°C
13. Ionization of nitric acid:  $\text{HNO}_3 \rightleftharpoons \text{NO}_3^- + \text{H}^+$   $K_a = 1.0 \times 10^7$  at 25°C
14. Ionization of perchloric acid:  $\text{HClO}_4 \rightleftharpoons \text{ClO}_4^- + \text{H}^+$   $K_a = 1.0 \times 10^7$  at 25°C
15. Ionization of sulfuric acid:  $\text{H}_2\text{SO}_4 \rightleftharpoons \text{HSO}_4^- + \text{H}^+$   $K_a = 1.0 \times 10^{-2}$  at 25°C
16. Ionization of hydrochloric acid:  $\text{HCl} \rightleftharpoons \text{Cl}^- + \text{H}^+$   $K_a = 1.0 \times 10^7$  at 25°C

# 19 2 Hydrogen Ions And Acidity Answer Key

**Terry C. Jones**



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