

Ch10 Study Guide

James

## Barriers

— 1 —

Chemists relate units of counting, of mass, and of volume to a single quantity called the 1. The number of representative particles in one mole of a substance is 2. To find the mass of a mole of a compound, scientists add together the 3 of the atoms in the compound. When you substitute the unit "grams" for amu, you obtain the 4 of the compound. There are 5 representative particles in a mole of any substance.

1 mole 2 Avogadro number 3 atomic mass 4 molar mass 5  $6.02 \times 10^{23}$

ATTENDS

ST. 6. a mole of a pure substance contains  $6.02 \times 10^{23}$  atoms.

Q. 7. The representative particle of a compound is the molecule.

At 8 A<sub>1</sub> mole of  $\text{CCl}_4$  is composed of one atom of carbon and 4 atoms of chlorine.

**AT** 9. A mole of carbon atoms has a mass approximately 3 times as great as the mass of a mole of helium atoms.

NT 10. The molar mass of nitrogen gas,  $\text{N}_2$ , is 14.02 g.

## Background

### B. 11. *Quercus ilex* L. (qv. 10)

subatomic molecules or ions in a substance.

### 6.13 nuclear fusion

1316 02 x 1023

13 module

c) the mass of 1 mole of a substance

### • 14 representative particles

④ SI unit that measures the amount of a substance

Find the molar mass of a)  $\text{NaNO}_3$  b)  $\text{Ca}(\text{HSO}_4)_2$  c)  $\text{CH}_3\text{NH}_2 \cdot \text{CO}_2$  d)  $\text{DN}_3\text{O}_2$

$$A = 22.99 + 14.91 + 16.99 + 16.99 + 16.99 = 85 \text{ g/mol}$$

B-52 00+52 00+52 00+32 00+32 00+16 00+16 00+16 00+16 00+16 00+16 00+16 00+16 00-348 128/mile

8.5e-001 mol 3.481.2e-001 mol 96.11e-001 mol 2.42.0e-001 mol

12

14.01+14.01+14.01+14.01+14.01+14.01+14.01+16.00+16.00+16.00+16.00+16.00+16.00+16.00+16.00+16.00=242.07g  
mixed

# Bju Chemistry Study Guide

**Theodore E. Brown,James C. Hill**

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