

Chemical Reactions:

1. The following reaction $\text{PbO}_2 + \text{H}_2\text{O}_2 \rightarrow \text{PbO}_2\text{O}_3 + \text{H}_2\text{O}$ is not true at 4.0 mol/l Pb²⁺:

- reactants with 2.0 mol/l Pb²⁺
- product excess with 3.0 mol/l Pb²⁺
- constant excess with 3.0 mol/l Pb²⁺
- product excess with 3.0 mol/l Pb²⁺
- reactants with 4.0 mol/l Pb²⁺

2. A 1.0 mol/l example of the NO_2 reaction is oxygenated according to the equation: $2\text{NO}_2\text{O}_2(\text{g}) \rightarrow 2\text{NO}_2\text{O}(\text{g}) + 2\text{O}_2(\text{g})$. Below, which mixture of O_2 and NO_2 will result in the same NO_2 concentration after 1000 s?

- 0.0 mol
- 1.0 mol
- 1.2 mol
- 0.7 mol
- 2.0 mol

3. The reaction of NO_2 is step characterized by the equation: $4\text{NO}_2 \rightarrow 3\text{O}_2 + \text{N}_2\text{O}_5$. If 0.0 mol/l NO_2 is 1.0 mol/l example of NO_2 , how many moles of O_2 will there be after the first 1000 s reaction time?

- 0.327 mol
- 0.790 mol
- 1.000 mol
- 1.34 mol
- 2.007 mol

4. Calculated the following reaction, at least 2000 s are necessary to calculate:

$$\text{NO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}_2\text{O}_3(\text{g}) + \text{NO}_2\text{O}(\text{g})$$

At 2.0 mol/l NO_2 and 2.0 mol/l O_2 initial reaction rate is 0.01 mol/l NO_2 / s. What is the final concentration of NO_2 ?

- 0.8
- 0.6
- 0.4
- 0.2
- 0.0

5. Under the following conditions, $\text{NO}_2 + \text{NO}_2(\text{g}) \rightarrow \text{NO}_2\text{O}(\text{g}) + \text{NO}(\text{g})$ is:

$$\text{NO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightarrow \text{NO}_2\text{O}(\text{g}) + \text{NO}(\text{g}) \rightleftharpoons \text{NO}_2\text{O}(\text{g}) + \text{NO}(\text{g})$$

What excess of oxygen, $\text{O}_2(\text{g})$ is required to react completely with 1.0 g of NO_2 ?

- 0.0001 x 10⁻² g
- 0.001 g
- 0.005 g
- 0.010 g
- 0.100 g

6. Under the following conditions, $\text{NO}_2 + \text{NO}_2(\text{g}) \rightarrow \text{NO}_2\text{O}(\text{g}) + \text{NO}(\text{g})$ is:

$$\text{NO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightarrow \text{NO}_2\text{O}(\text{g}) + \text{NO}(\text{g}) \rightleftharpoons \text{NO}_2\text{O}(\text{g}) + \text{NO}(\text{g})$$

What excess of oxygen, $\text{O}_2(\text{g})$ is required to react completely with 1.0 g of NO_2 ?

- 0.0001 x 10⁻² g
- 0.001 g
- 0.005 g
- 0.010 g
- 0.100 g

7. Calculated the following reaction, at least 2000 s are necessary to calculate:

$$\text{NO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{NO}_2\text{O}(\text{g}) + \text{NO}_2\text{O}(\text{g})$$

At 2.0 mol/l NO_2 and 2.0 mol/l O_2 initial reaction rate is 0.01 mol/l NO_2 / s. What is the final concentration of NO_2 ?

- 0.790 g NO_2
- 1.000 g NO_2
- 0.901 g NO_2
- 1.200 g NO_2
- 1.400 g NO_2

8. Under the following conditions, $\text{NO}_2 + \text{NO}_2(\text{g}) \rightarrow \text{NO}_2\text{O}(\text{g}) + \text{NO}(\text{g})$ is:

$$\text{NO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightarrow \text{NO}_2\text{O}(\text{g}) + \text{NO}(\text{g}) \rightleftharpoons \text{NO}_2\text{O}(\text{g}) + \text{NO}(\text{g})$$

What excess of oxygen, $\text{O}_2(\text{g})$ is required to react completely with 1.0 g of sulfur in the following reaction:

$$\text{S}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{SO}_2(\text{g})$$

- 0.0001 x 10⁻² mol/mol/mol O_2
- 0.001 g O_2
- 0.01 g O_2
- 0.020 g O_2
- 0.100 g O_2

9. The following reaction $\text{2NO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2\text{O}(\text{g})$ is:

Under the same conditions of concentration, which excess of oxygen, $\text{O}_2(\text{g})$ would the product excess? 1.000 mol/l NO_2 and 0.0001 mol/l O_2 ?

- 1.000 mol/l
- 0.0001 mol/l
- 0.00001 mol/l
- 0.000001 mol/l

10. Under the following conditions, $\text{2NO}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2\text{O}(\text{g})$ is:

What excess of oxygen, $\text{O}_2(\text{g})$ is required to react completely with 1.0 g of sulfur in the following reaction:

$$\text{S}(\text{s}) + \text{O}_2(\text{g}) \rightarrow \text{SO}_2(\text{g})$$

- 0.0001 x 10⁻² mol/mol/mol O_2
- 0.001 g O_2
- 0.01 g O_2
- 0.020 g O_2
- 0.100 g O_2

11. Under the same conditions of oxygen, $\text{O}_2(\text{g})$, required to react with 1.0 g of sulfur in the following reaction:

$$\text{2NO}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{NO}_2\text{O}(\text{g})$$

- 0.0001 x 10⁻² mol/mol/mol O_2
- 0.001 g O_2
- 0.01 g O_2
- 0.020 g O_2
- 0.100 g O_2

Multiple Choice Stoichiometry Test

Ardas Ozsogomonyan

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Cambridge Checkpoints VCE Chemistry Units 3 and 4 2013 Roger Slade, Maureen Slade, 2012-12-14 Cambridge Checkpoints VCE are updated regularly to provide you with the most up to date exam preparation available **Cracking the AP Chemistry Exam, 2012 Edition** Paul Foglino, 2011-08-02 Provides techniques for achieving high scores on the AP chemistry exam and includes two full length practice tests *Roadmap to the Regents* Sasha Alcott, 2003 If Students Need to Know It It's in This Book This book develops the chemistry skills of high school students It builds skills that will help them succeed in school and on the New York Regents Exams Why The Princeton Review We have more than twenty years of experience helping students master the skills needed to excel on standardized tests Each year we help more than 2 million students score higher and earn better grades We Know the New York Regents Exams Our experts at The Princeton Review have analyzed the New York Regents Exams and this book provides the most up to date thoroughly researched practice possible We break down the test into individual skills to familiarize students with the test's structure while increasing their overall skill level We Get Results We know what it takes to succeed in the classroom and on tests This book includes strategies that are proven to improve student performance We provide a breakdown of the skills based on New York standards and objectives hundreds of practice questions organized by skill two complete practice New York Regents Exams in Physical Setting Chemistry *Master the CLEP* Peterson's, 2012-04-30 Peterson's Master the CLEP is the essential prep guide for non traditional students who are looking to begin a college career and for current students who want to save on tuition costs by testing out of certain courses When you use this guide you will get a basic overview of the CLEP a look at the general examinations and the vast number of subject specific exams tips for registering and preparing for the CLEP and a guide to interpreting your test scores This is followed by extensive reviews of the five CLEP general examinations with pre and post tests and subject matter overviews that will help you pinpoint your strengths and weaknesses The Appendix will go into detail about a number of CLEP subject exams *CLEP Success* Peterson's, 2007-05-04 Offers advice on improving scores on the CLEP and includes reviews for five CLEP subjects and five full length practice tests with explanatory answers

Journal of Chemical Education, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers *Tests* Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services

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provide a higher quality of learning for students Optimizing STEM Education With Advanced ICTs and Simulations is an innovative reference source for the latest scholarly research on the integration of digital tools for enhanced STEM based learning environments Highlighting a range of pivotal topics such as mobile games virtual labs and participatory simulations this publication is ideally designed for educators professionals academics and students seeking material on emerging educational technologies

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A Summary of Science Education, 1988 Leopold E. Klopfer,1990-06 **Three Cognitive Skills in Chemistry and Their Application to Stoichiometry** Ardas Ozsogomonyan,1977

Cracking the AP Chemistry, 2002-2003 Edition Paul Foglino,Princeton Review (Firm),2002-01-29 The Princeton Review realizes that acing the AP Chemistry exam is very different from getting straight As in school The Princeton Review doesn t try to teach students everything there is to know about chemistry only the techniques they ll need to score higher on the exam There s a big difference In Cracking the AP Chemistry TPR will teach test takers how to think like the test makers and Score higher by reviewing key chemistry concepts Earn more points by becoming familiar with the format of the test Safeguard against traps that can lower scores Perfect skills with review questions in each chapter This book includes 2 full length simulated AP Chemistry exams All of The Princeton Review practice test questions are like the ones test takers will see on the actual exam and every solution is fully explained

Computer-assisted Instruction in Chemistry: General approach James S. Mattson,Harry B. Mark,Hubert C. MacDonald,1974 **Proceedings of Conference on Computers in the Undergraduate Curricula ,1973**

Media Review ,1986 **Nigerian Journal of Research and Production ,2010-11** **Science Instruction in the Middle and Secondary Schools** Eugene L. Chiappetta,Thomas R. Koballa (Jr.),Alfred T. Collette,1998

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