

4. a. A sample of a $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}/\text{K}_2\text{C}_2\text{O}_4 \cdot \text{H}_2\text{O}$ solid salt mixture is dissolved in ~150 ml. of deionized water until it reaches a pH that is basic. The precipitate, after having been filtered, was air-dried and weighed. Data for the first trial are shown below. Complete the following table. (See Report Sheet.) Record calculated values with the correct number of significant figures.

A. Precipitation of $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ from the Salt Mixture

1. Mass of salt mixture (g)
2. Mass of filter paper (g)
3. Mass of filter paper and $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ (g)
4. Mass of air-dried $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ (g)

<i>Calculation Zone</i>	
Data Analysis, 1.	
0.879	
1.896	
2.180	
<u>0.231</u>	

B. Determination of Limiting Reactant

1. Limiting reactant in salt mixture
2. Excess reactant in salt mixture



Data Analysis

1. Moles of $\text{CaC}_2\text{O}_4 \cdot \text{H}_2\text{O}$ precipitated (mol)
Show calculation.
2. Moles of limiting reactant in salt mixture (mol)
See equation B.1.
3. Mass of limiting reactant in salt mixture (g)
Show calculation.
4. Mass of excess reactant in salt mixture (g)
Excess mass of salt mixture minus mass of limiting reactant.
Show calculation.
5. Percent limiting reactant in salt mixture (%)
Show calculation.
6. Percent excess reactant in salt mixture (%)

Data Analysis, 4.

0.000100

0.000100

0.000100

0.000100

Data Analysis, 5.

0.000100

0.000100

4. b. For Trials 2 and 3, the percent $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ in the mixture was 39.7% and 28.4% respectively.
a. What is the average percent of $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ in the sample? Data Analysis, B.

b. What are the standard deviation and the relative standard deviation (% RSD) for the percent of $\text{CaCl}_2 \cdot 2\text{H}_2\text{O}$ in the sample? Data Analysis, C and D.

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