

Acids & Bases HOMEWORK

Lesson 1

1. Provide the name for each of the strong acids listed below. Provide the formula for each of the strong bases listed below.

- HCl **hydrochloric acid**
- H₂SO₄ **sulfuric acid**
- HNO₃ **nitric acid**
- HClO₄ **perchloric acid**
- Strontium hydroxide **LiOH**
- potassium hydroxide **KOH**
- calcium hydroxide **Ca(OH)₂**
- barium hydroxide **Ca(OH)₂**
- ammonium hydroxide **Li(OH)₄**

2. Of the acids and bases listed below, circle the ones that completely dissociate in water.
HClO₄ is STRONG with a lower cell completely dissociate in water

- HCl** **H₂** **H₂O** **Mg(OH)₂** **HNO₃** **Ca(OH)₂** **HClO₄** **H₂SO₄** **H₂CO₃**

3. What is the formula for hydrosulfuric acid? **H₂S**

4. H₂CO₃ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for the reaction. **H₂CO₃ + H₂O → H₂O⁺ + HCO₃⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because H₂CO₃ is a weak acid.**
- Would you define H₂CO₃ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is H₂CO₃ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of H₂CO₃. **HCO₃⁻**
- Identify H₂CO₃ as polyprotic, amphoteric, or neither. **polyprotic, because it contains multiple H's to give.**

5. F⁻ will react with water and gain a H⁺ ion in the process.

- Write a balanced equation for the reaction. **F⁻ + H₂O → HF + OH⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because F⁻ is a weak base.**
- Would you define F⁻ as a proton donor or a proton acceptor? **proton acceptor, because it is gaining a H⁺ ion.**
- In your written equation, is F⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry base, because it is getting a proton.**
- Identify the conjugate acid of F⁻. **HF**
- Identify F⁻ as polyprotic, amphoteric, or neither. **neither**

6. HSO₄⁻ will react with water and lose a H⁺ ion in the process.

- Write a balanced equation for the reaction. **HSO₄⁻ + H₂O → H₂O⁺ + SO₄²⁻**
- Would you expect a large or small equilibrium constant for this reaction? **small, because HSO₄⁻ is a weak acid.**
- Would you define HSO₄⁻ as a proton donor or a proton acceptor? **proton donor, because it is losing a H⁺ ion.**
- In your written equation, is HSO₄⁻ acting as a Brønsted-Lowry acid or a Brønsted-Lowry base? **Brønsted-Lowry acid, because it is giving a proton.**
- Identify the conjugate base of HSO₄⁻. **SO₄²⁻**
- Identify HSO₄⁻ as polyprotic, amphoteric, or neither. **amphoteric, because of both a negative charge (only will take a proton) and a H to give (can give out H's too).**

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