

## Stoichiometry: Multiple Choice

1. The limiting reactant of a chemical reaction is:

- A) the reactant of a chemical process that is not consumed completely
- B) the product of a chemical process that is consumed completely
- C) the product of a chemical process that is not consumed completely
- D) the reactant of a chemical process that is consumed completely
- E) none of the above

ANS: D

2. Consider the following reaction:  $\text{Ca}(\text{OH})_{2(\text{aq})} + \text{Al}_2(\text{SO}_4)_{3(\text{aq})} \rightarrow \text{CaSO}_4(\text{aq}) + \text{Al}(\text{OH})_{3(\text{aq})}$

When balanced, the mole ratio for calcium hydroxide reacting with aluminum sulphate is

- A) 1:1
- B) 2:1
- C) 2:3
- D) 1:2
- E) 3:1

ANS: E

3. Sodium phosphate,  $\text{Na}_3\text{PO}_4$ , can be prepared by combining sodium hydroxide and hydrogen phosphate as shown in the following balanced chemical equation:



If 2.5 mol of  $\text{H}_2\text{PO}_4$  are reacted, the number of moles of water that will be produced is:

- A) 1.0
- B) 2.5
- C) 3.0
- D) 5.0
- E) 7.5

ANS: E

4. Potassium chloride and oxygen gas are formed upon heating potassium chlorate. The following balanced equation shows this decomposition process:



The number of moles of  $\text{KClO}_3$  required to produce 5.0 mol of oxygen is:

- A) 1.0
- B) 2.0
- C) 3.0
- D) 3.3
- E) 5.0

ANS: D

5. Determine the number of grams of  $\text{CO}_2$  that can be produced from 48.0 g of the hydrocarbon  $\text{C}_2\text{H}_6$ , according to the following equation:



- A) 37.6 g
- B) 50.2 g
- C) 87.8 g
- D) 154 g
- E) 188.4 g

6. The limiting reactant in a chemical reaction is the \_\_\_\_\_ that is \_\_\_\_\_.

- A) Reactant, in excess
- B) Product, left over
- C) Reactant, used up
- D) Product, used up

ANS: C

# Multiple Choice Stoichiometry Test

**Paul Foglino, Princeton Review (Firm)**

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**Journal of Chemical Education**, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers *Tests* Richard C. Sweetland, Daniel J. Keyser, 1983 Descriptions of over 3000 tests in English intended as a guide for psychologists educators and other personnel who need test information to meet their assessment needs Topical arrangement under 3 main sections of psychology education and business Each entry gives test name associated personal names intended age group purpose description time range scoring cost and publisher Indexes by test titles authors publishers visually impaired tests and scoring services

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**A Summary of Science Education, 1988** Leopold E. Klopfer,1990-06 **Three Cognitive Skills in Chemistry and Their Application to Stoichiometry** Ardas Ozsogomonyan,1977

**Cracking the AP Chemistry, 2002-2003 Edition** Paul Foglino,Princeton Review (Firm),2002-01-29 The Princeton Review realizes that acing the AP Chemistry exam is very different from getting straight As in school The Princeton Review doesn t try to teach students everything there is to know about chemistry only the techniques they ll need to score higher on the exam There s a big difference In Cracking the AP Chemistry TPR will teach test takers how to think like the test makers and Score higher by reviewing key chemistry concepts Earn more points by becoming familiar with the format of the test Safeguard against traps that can lower scores Perfect skills with review questions in each chapter This book includes 2 full length simulated AP Chemistry exams All of The Princeton Review practice test questions are like the ones test takers will see on the actual exam and every solution is fully explained

**Computer-assisted Instruction in Chemistry: General approach** James S. Mattson,Harry B. Mark,Hubert C. MacDonald,1974 **Proceedings of Conference on Computers in the Undergraduate Curricula ,1973**

**Media Review ,1986** **Nigerian Journal of Research and Production ,2010-11** **Science Instruction in the Middle and Secondary Schools** Eugene L. Chiappetta,Thomas R. Koballa (Jr.),Alfred T. Collette,1998

New edition of a text providing far more than simply a cookbook of activities for science teachers Coverage includes discussion on the nature of science national standards and innovative programs the nature of adolescent learners and their schools teaching strategies and classroom management pl *The Education Index* ,1984

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### **Table of Contents Multiple Choice Stoichiometry Test**

1. Understanding the eBook Multiple Choice Stoichiometry Test
  - The Rise of Digital Reading Multiple Choice Stoichiometry Test
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Stoichiometry Test
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Stoichiometry Test
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Stoichiometry Test

- Personalized Recommendations
- Multiple Choice Stoichiometry Test User Reviews and Ratings
- Multiple Choice Stoichiometry Test and Bestseller Lists

5. Accessing Multiple Choice Stoichiometry Test Free and Paid eBooks

- Multiple Choice Stoichiometry Test Public Domain eBooks
- Multiple Choice Stoichiometry Test eBook Subscription Services
- Multiple Choice Stoichiometry Test Budget-Friendly Options

6. Navigating Multiple Choice Stoichiometry Test eBook Formats

- ePub, PDF, MOBI, and More
- Multiple Choice Stoichiometry Test Compatibility with Devices
- Multiple Choice Stoichiometry Test Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Multiple Choice Stoichiometry Test
- Highlighting and Note-Taking Multiple Choice Stoichiometry Test
- Interactive Elements Multiple Choice Stoichiometry Test

8. Staying Engaged with Multiple Choice Stoichiometry Test

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Multiple Choice Stoichiometry Test

9. Balancing eBooks and Physical Books Multiple Choice Stoichiometry Test

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Multiple Choice Stoichiometry Test

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Multiple Choice Stoichiometry Test

- Setting Reading Goals Multiple Choice Stoichiometry Test
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Multiple Choice Stoichiometry Test

- Fact-Checking eBook Content of Multiple Choice Stoichiometry Test
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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