

Chemistry 2nd Semester Exam Review Sheet:**Chapters 9, 10, 11, 12, 13, 14, 15, and 16****Chapter 9: Chemical Quantities**

1. Solve stoichiometry problems using grams \rightarrow moles \rightarrow moles \rightarrow grams calculations. Examples below:
a. Calculate moles/grams of $(\text{NH}_4)_2\text{SO}_4$ produced when 30.0 moles of NH_3 react with H_2SO_4 .



$$30.0 \text{ moles} \times 1 \text{ mol}/2 \text{ mol} = 15.0 \text{ moles}$$

$$15.0 \text{ moles } (\text{NH}_4)_2\text{SO}_4 \times 132.2 \text{ g/mol/mole} = 1983 = 1980 \text{ grams (to 3 s.f.)}$$

b. $\text{Na}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{NaNO}_3 + \text{CO}_2 + \text{H}_2\text{O}$. How many moles/grams of sodium nitrate are produced by the complete reaction of 100. grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? *Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.*

Grams \rightarrow Moles \rightarrow Moles \rightarrow Grams

$$100. \text{ g } \text{Na}_2\text{CO}_3 \times 1 \text{ mol}/106.0 \text{ g} \times 2 \text{ mol } \text{NaNO}_3 / 1 \text{ mol } \text{Na}_2\text{CO}_3 \times 85.00 \text{ g/mol} = 160.4 \text{ g}$$

$$500.0 \text{ g } \text{NaNO}_3 \times 1 \text{ mol}/85.00 \text{ g} \times 2 \text{ mol} / 2 \text{ mol} \times 63.02 \text{ g/mol} = 370.7 \text{ grams}$$

2. Calculate percent yield of products. Example below using (b) above.

What is the theoretical yield of NaNO_3 if 100.0 grams of Na_2CO_3 are used in the reaction? What is the percent yield of NaNO_3 if 130.5 grams of NaNO_3 are actually collected?

Theoretical yield = 160.4 grams (answer from above)

$$\text{Percent Yield} = (\text{Actual Yield/Theoretical Yield}) \times 100 = (130.5 \text{ g}/160.4 \text{ g}) \times 100 = 81.36\%$$

3. What is a limiting reactant in a chemical reaction? The reactant that is used up first, thereby limiting the amount of product that can be formed.

Chapter 10: Modern Atomic Theory

4. Describe Rutherford's model of the atom. What could it not explain?

Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.

5. Define electromagnetic radiation, wavelength, and frequency.

6. How do atoms gain energy to become excited? Absorb a photon of energy.

How do they lose energy to return to their ground state? Release a photon of same energy as that absorbed.

7. What is the evidence that energy absorbed by the hydrogen atom is quantized? Specific wavelengths of light seen in emission spectra for atoms (page 285-86)

8. Describe the Bohr model of the atom. How did it improve on Rutherford's model?

Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

See book, notes, probs 54, 96

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New Trends in Chemistry Teaching Unesco,1981 The American Secondary School Leslie Owen Taylor,1960

Teaching Classics Worldwide Steven Hunt,John Bulwer,2025-02-06 Designed for instructors in schools and universities as well as stakeholders and policy makers in education everywhere this book is a systematic guide to contemporary school teaching of classical languages literature and civilisation in major countries across the world This book examines the trends and assesses the rationale for such diversity in programmes of study and asks the question What are classical subjects for in today s schools Each chapter is arranged by geographical area and draws on the experiences of teachers and other education experts in each country commenting on contemporary practices Themes such as the impact of national education policies enrolment assessment accessibility and inclusion are discussed The teaching of Classics has wide variations in practices and purposes from one education system to another in some the teaching of Classics continues to be focused on the mastery of language and grammar in others active use of the ancient languages is growing in popularity and yet in others the focus is shifting towards the study of the languages and civilisations in translation Presenting an overview survey and comparison of practices across the world this book is essential reading for instructors and teaching training courses in the world of classical education *Resources in education* ,1987-09 **Cornell University Courses of Study** Cornell University,2003

Catalog Issue University of Colorado,1974 **Summer School Number** Southwestern at Memphis,1933 **Source**

Book for Chemistry Teachers William Thomas Lippincott,1981 **American Journal of Physics** ,1994 Who's who in America John William Leonard,Albert Nelson Marquis,1901 Vols 28 30 accompanied by separately published parts with title Indices and necrology **Abridged Index Medicus** ,1976 **Computer Aided Training in Science and Technology** E. Oñate,1990 **Resources in Education** ,1978 Paperbound Books in Print ,1966 **University of Colorado at Denver Catalog** University of Colorado at Denver,1992 **Current Index to Journals in Education, Semi-Annual Cumulation, July-December, 1976** ,1976-10 **Computer-assisted Instruction** Albert E. Hickey,1968 **Current Index to Journals in Education** ,1979 **Science Books & Films** ,1980 Documentation Abstracts ,1969

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