



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base
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Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

$$\text{pH} = -\log_{10}[\text{H}^+]$$

Practice Calculations

Use the following equations to calculate the pH of a solution with a given concentration of hydrogen ions.

Hydrogen ion concentration: H^+ concentration (hydrogen ions) \times concentration of other ions \times concentration of water \times concentration of other ions

Hydrogen ion concentration: H^+ concentration (hydrogen ions) \times concentration of other ions \times concentration of water \times concentration of other ions \times concentration of other ions

Hydrogen ion concentration: H^+ concentration (hydrogen ions) \times concentration of other ions \times concentration of water \times concentration of other ions \times concentration of other ions

Answers

1. $\text{pH} = -\log_{10}[\text{H}^+]$	2. $\text{pH} = \log_{10}[\text{OH}^-]$
3. $\text{pH} = -\log_{10}[\text{H}^+]$	4. $\text{pH} = \log_{10}[\text{OH}^-]$
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99. $\text{pH} = -\log_{10}[\text{H}^+]$	100. $\text{pH} = \log_{10}[\text{OH}^-]$

19 2 Hydrogen Ions And Acidity Work Sheet Answers

S Ben Porath

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