

Acids, Bases & pH Review - Honors Chemistry

Directions: For all multiple choice questions, highlight the correct answer(s). Answer short answer questions in the box provided.

Part 1: Properties & Reactions of Acids & Bases

1. Complete the table below.

	ACIDS	BASES
Table		
pH range		
Conducting electricity		
Potassium		
Lithium		
Phenolphthalein		

2. An indicator:

- a. is the type of acid.
- b. is the type of base.
- c. has a neutral pH.
- d. changes color depending on the pH.

3. Lithium and phenolphthalein are indicators. What does this mean?

4. Acids and bases "dissociate" in water to form H^+ (or H_3O^+) and OH^- (or H_2O_2^-). What does "dissociate" mean in this context?

5. Identify whether each substance is an acid or a base.

- a. H_2O^+ :
- b. HCl (HCl):
- c. $\text{Al(OH}_3\text{)}_3$:
- d. $\text{NH}_3 + \text{H}_2\text{O} \rightarrow \text{NH}_4^+ + \text{OH}^-$; NH_3 is weak.
- e. $\text{H}_2\text{SO}_4 + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{HSO}_4^-$; H_2SO_4 is strong.
- f. $\text{HNO}_3 \rightarrow \text{H}^+ + \text{NO}_3^-$; HNO_3 is strong.

Acid And Base Study Guide 2011

Nehemiah Hawkins

Acid And Base Study Guide 2011:

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