

- Structure, Hybridization, Resonance, Aromaticity
 - # of bonds-1= sp (e.g. 3 bonds = sp^2)
 - Aromaticity= $4n+2$ electrons
 - Strained ring causes greater heat of combustion
- Acids and Bases
 - Lewis acid= accepts electron pair
 - Higher percentage of s character means more acidic hydrogen
- Stereoisomerism
 - Constitutional Isomers
 - Isomers that have different atom connectivity
 - Stereoisomers
 - Same connectivity but different arrangement in space
 - Enantiomers
 - Stereoisomers that are non-superimposable mirror reflections
 - Diastereomers
 - Stereoisomers that are not mirror reflections of each other
 - Stereocenter= sp^3 carbon bonded to four different substituents
 - Assigning R/S configuration
 - If lowest group is pointing away, determine normally
 - If lowest group is not pointing away, exchange with group pointing away and determine reverse

- Nucleophilic Substitutions and Eliminations

| | | | |
|----------------|---------------|--------------------------------------|---------------------------------------|
| ◦ Alkyl Halide | ◦ Strong Base | ◦ Weak Base | ◦ Poor Nucleophile |
| ◦ 1 | ◦ S_N2 | ◦ S_N2 | ◦ No rxn. |
| ◦ 2 | ◦ Mostly E2 | ◦ S_N2 | ◦ $S_N1/E1$ in polar, protic solvents |
| ◦ 3 | ◦ E2 | ◦ $S_N1/E1$ in polar, protic solvent | ◦ $S_N1/E1$ in polar, protic solvents |

- Relative strength of nucleophiles
 - Negatively charged nucleophiles react faster in substitution reactions than neutral nucleophiles
 - Nucleophilic strength is in the same order as base strength when the nucleophilic atoms are in the same period of the periodic table
 - Nucleophilic strength increases with increasing atom size when the nucleophilic atoms are in the same column of the periodic table
- Leaving group ability
 - Increases with decreasing base strength
- Base strength and structure

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