

Acids & Bases: Assessment Review

Name _____

1. **True/False** – Write True if the statement is true, False if the statement is false, Conjugate Acid and Conjugate Base if an acid and a base are mixed together.

2. Rewrite a chemical equation showing how HClO_4 behaves as an acid when dissolved in water. Label the acid, base, conjugate acid & conjugate base.

3. Rewrite a chemical equation showing how ammonia (NH_3) behaves as a base when dissolved in water. Label the acid, base, conjugate acid & conjugate base.

4. **True/False** – Write True if the calculation with either $[\text{H}^+]$ or $[\text{OH}^-]$ shows that H_2O is a neutral acid or pH shows pH < 7 (a neutral pH check calculation.)

5. Recreate the calculation flowchart on the diagram provided. (be careful with these numbers!)



6. Calculate $[\text{OH}^-]$ in a solution that $[\text{H}^+] = 7.2 \times 10^{-7}$ M. Is the solution acidic, basic, or neutral?

7. Calculate $[\text{H}^+]$ in a solution of $[\text{OH}^-] = 5.2 \times 10^{-11}$ M. Is the solution acidic, basic, or neutral?

8. Calculate the $[\text{H}^+]$ in a solution that shows a pH of 1.1. Is the solution acidic, basic, or neutral?

9. Calculate the $[\text{OH}^-]$ in a solution of $[\text{OH}^-] = 1.2 \times 10^{-9}$ M if the solution acidic, basic, or neutral?

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William Albert Noyes

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