

Acids & Bases: Assessment Review

Name _____

Skill – The student will be able to identify the acid, base, conjugate acid, and conjugate base in an acid – base reaction.

1. Write a chemical equation showing how HCO_3^- behaves as an acid when dissolved in water. Label the acid, base, conjugate acid & conjugate base.

2. Write a chemical equation showing how ammonia (NH_3) behaves as a base when dissolved in water. Label the acid, base, conjugate acid & conjugate base.

Skill – The student will be able to calculate with the $[\text{H}^+]$ or $[\text{OH}^-]$ from the ion product (K_w) or pH from pOH (a one-step pH block calculation).

3. Recreate the calculation flowchart on the diagram provided. (Be careful...it's been flipped!)



4. What is $[\text{OH}^-]$ if a solution has $[\text{H}^+] = 7.1 \times 10^{-4} \text{ M}$? Is the solution acidic, basic, or neutral?

5. Calculate $[\text{H}^+]$ in a solution if $[\text{OH}^-] = 5.1 \times 10^{-11} \text{ M}$. Is the solution acidic, basic, or neutral?

6. Calculate the $[\text{H}^+]$ in a solution that shows a pOH of 11.75. Is the solution acidic, basic, or neutral?

7. What is the pOH of a solution if $[\text{OH}^-] = 5.0 \times 10^{-4} \text{ M}$? Is the solution acidic, basic, or neutral?

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William Albert Noyes



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