

Chemistry: Acids and Bases Study Guide (Ch. 15)

1. Give the definitions of an acid and a base according to the Arrhenius' Definition.
2. Give the definition of an acid and a base according to the bronsted-lowry definition.
3. Which of the definitions is used more often? Why is this?
4. What two substances are created when an acid is mixed with a base?
5. What are the tastes of acids and bases?
6. How does a base feel? How does that compare to the texture of acids?
7. What type of gas is released when an acid mixes with a metal (assuming that there is a reaction)?
8. Acids and bases break down into charge ions called _____ that can conduct electricity. Does strong or weak acid/base conduct more electricity?
9. What is an indicator? What are the 3 types of indicators in the form of paper? What are some indicators in the form of solution?
10. What indicator solution is clear in an acid, but turns pink in a base (pH >8)?
11. What is the ion product (number) for water at 25 °C?
12. What is an amphoteric substance?
13. What determines the strength of an acid? List strong acids and weak acids.
14. How is the strength of a base determined? List strong bases and weak bases.
15. Give the formulas of the following acids: hydrochloric acid, nitrous acid, acetic acid, carbonic acid, and sulfuric acid.
16. What are the names of the following acids? HNO_3 , HF , H_3PO_4 , H_2SO_4 .
17. What color does blue litmus paper turn in an acidic solution? A basic solution?
18. What is the range of pH for an acid or a base of the known concentration?
19. What is a solution with a pH=7 called? Give an example of a solution with this pH.
20. Describe the concept of conjugate acid-base pair. What is the different between acid and conjugate acid? Base and conjugate base? Show an example.
21. What is a buffer solution?
22. What type of solution contains an acid or base of a known concentration in titration?
23. What is the pH of a solution where $[\text{OH}^-]$ is higher than $[\text{H}_3\text{O}^+]$? Give a range.
24. If the $[\text{H}_3\text{O}^+]$ is greater than $[\text{OH}^-]$, is the solution acidic or basic?
25. What is the end point of titration? How is this different from equivalence point?
26. What should the end point of titration be close to?
27. What is a buffer capacity?
28. IF $[\text{H}_3\text{O}^+]$ or $[\text{OH}^-]$ are added beyond the buffer capacity, what will happen to the pH?

Problems

29. If pH is 4, what is pOH?
30. In a solution of water, the concentration of OH^- ions is $1.0 \times 10^{-6}\text{M}$. What is the concentration of the hydronium ions?
31. The concentration of OH^- ions in a solution is $3.4 \times 10^{-3}\text{M}$. Is the solution acidic basic or neutral? What is the pH of the solution?
32. A volume of 37 mL of 0.36 M KCN neutralizes a 75-mL sample of HClO solution (acid). What is the concentration of HClO ?

Acid And Bases Study Guide Ch 15

Paul A. Bartlett, Judith G. Koch



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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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General Chemistry Study Guide Sixth Edition Darrell D. Ebbing, 1999

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Study Guide, Introductory Chemistry, a Foundation ; Introductory Chemistry ; Basic Chemistry Iris Stovall, Donald J. DeCoste, 2000

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Study Guide to Accompany Chemical Principles, Properties, and Reactions Kenneth L. Henold, 1984

Student Study Guide to Accompany Petrucci's General Chemistry Robert K. Wismer, 1985

Study Guide with Selected Solutions Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems

Ess Chem Probs Study Guide David Margolese, David W. Oxtoby, 2004-04

Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application

Student Study Guide to Accompany Chemistry Kenneth J. Hughes, Paul B. Kelter, 2002-05 Reviews essential points from the text presents problems and hints to assist the student in solving them Solutions to selected problems in the end of chapter problem sets are included By Kenneth Hughes of the University of Wisconsin Madison

Study Guide to Organic Chemistry, 4th Ed

Robert Thornton Morrison, Robert Neilson Boyd, 1983 A popular introduction to organic chemistry which stresses the importance of molecular structure in understanding the properties and principles of organic chemistry Provides a wide variety of spectra to be analyzed Features four color photographs throughout **Study Guide to Accompany Principles of Genetics, 3rd Edition** D. Peter Snustad, Harold James Price, Michael J. Simmons, 2002-09-09 High quality illustrations with stepped out art to help readers visualize complex processes Human genetics and the role of the geneticist highlighted throughout Two new features in each chapter introductory Key Questions and closing Basic Exercises College Chemistry Peter C. Scott, 1980

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