

ACIDS, BASES AND SALTS

Definition of an acid An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	What causes the acidic properties of acids? The hydrogen ions (H^+) cause the acidic properties and these are formed in the presence of water.
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Another term to refer to an acid An acid is called a proton donor	Why an acid is also called a proton donor? It's because an acid provides protons or hydrogen ions (H^+) to other substances during the reaction.
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Substance to which an acid provides protons BASE	A base is then called PROTON ACCEPTOR	Why the base is called a proton acceptor? Because the base accepts hydrogen ions from acids	Equation for the reaction between acid and base $H^{+}_{(aq)} + OH^{-}_{(aq)} \longrightarrow H_2O_{(l)}$
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Common laboratory acids		These three common laboratory acids are also called MINERAL ACIDS	Why these acids are also called mineral acids? They are derived from mineral salts ie chlorides for HCl , sulphates for H_2SO_4 and nitrates for HNO_3
Hydrochloric acid	HCl		
Sulphuric acid	H_2SO_4		
Nitric acid	HNO_3		

Other mineral acids known	Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid	H_2SO_3	Derived from sulphites	CITRIC ACID from lemons
Carbonic acid	H_2CO_3	Derived from carbonates	TARTARIC ACID from grapes
Phosphoric acid	H_3PO_4	Derived from phosphates	ACETIC ACID from vinegar
Nitrous acid	HNO_2	Derived from nitrates	LACTIC ACID from sour milk
		Methanoic acid ($HCOOH$)	Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces HYDROGEN IONS	Term given to the number of hydrogen ions produced by one molecule of an acid BASICITY OF AN ACID	Definition of basicity of an acid BASICITY of an acid is the <u>number of hydrogen ions produced by one molecule of an acid in aqueous solution</u> .
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Acid Bases Study Guide

Leo J. Malone

Acid Bases Study Guide:

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1. Understanding the eBook Acid Bases Study Guide
 - The Rise of Digital Reading Acid Bases Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid Bases Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid Bases Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid Bases Study Guide
 - Personalized Recommendations
 - Acid Bases Study Guide User Reviews and Ratings
 - Acid Bases Study Guide and Bestseller Lists
5. Accessing Acid Bases Study Guide Free and Paid eBooks
 - Acid Bases Study Guide Public Domain eBooks
 - Acid Bases Study Guide eBook Subscription Services
 - Acid Bases Study Guide Budget-Friendly Options
6. Navigating Acid Bases Study Guide eBook Formats

- ePub, PDF, MOBI, and More
- Acid Bases Study Guide Compatibility with Devices
- Acid Bases Study Guide Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Acid Bases Study Guide
- Highlighting and Note-Taking Acid Bases Study Guide
- Interactive Elements Acid Bases Study Guide

8. Staying Engaged with Acid Bases Study Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Acid Bases Study Guide

9. Balancing eBooks and Physical Books Acid Bases Study Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Acid Bases Study Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Acid Bases Study Guide

- Setting Reading Goals Acid Bases Study Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Acid Bases Study Guide

- Fact-Checking eBook Content of Acid Bases Study Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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