

Geometry Final Spring 2021 Form A

1. Expand or Multiply (FOIL):  $(2x - 7)(x + 3)$

A.  $2x^2 - x - 21$

B.  $2x^2 + x - 21$

C.  $2x^2 - 21$

D.  $-42x^2$

Commented (P1): To turn off these comments:  
- Select the "Notes" tab.  
- On "Sharing," select "Final" instead of "Final Drawing Mode."

2. Consider the triangle shown below. What is the most appropriate equation to solve for the unknown side marked  $b$  in the figure?

A.  $b^2 = 19^2 + 15^2$

B.  $b^2 = 15^2 + 19^2 - 2(15)(19)\cos 73^\circ$

C.  $\frac{b}{\sin 73^\circ} = \frac{19}{\cos 73^\circ}$

D.  $\sin 73^\circ = \frac{b}{19}$



Commented (P2): A

Commented (User2): B

3. Find the value of  $d$ .

A.  $16$



Commented (P4): A

B.  $60$

C.  $1.58$

D.  $79$

4. Find the total surface area of the square-based pyramid.

A.  $12000 \text{ in}^2$



Commented (P5): A

B.  $2000 \text{ in}^2$

C.  $816 \text{ in}^2$

D.  $456 \text{ in}^2$

5. Find the equation of the circle with center  $(3, -4)$  and radius 6.

A.  $(x - 3)^2 + (y + 4)^2 = 6$

B.  $(x + 3)^2 + (y - 4)^2 = 6$

C.  $(x - 3)^2 + (y + 4)^2 = 36$

D.  $(x + 3)^2 + (y - 4)^2 = 36$

Commented (P6): C

6. Simplify:  $\sqrt{75}$

A.  $3\sqrt{5}$

B.  $5\sqrt{3}$

C.  $25\sqrt{3}$

D. It is already simplified.

Commented (P7): A

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web what is the percent composition of cu in terms of these two isotopes answer 69 15 cu 63 and 30 85 cu 65 figure pageindex 5 isotopes of an element are atoms with the same atomic number but different mass numbers isotopes of an element therefore differ from each other only in the number of neutrons within the nucleus

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web jan 11 2013 isotopes of an element share the same number of protons but have different numbers of neutrons let s use carbon as an example there are three isotopes of carbon found in nature carbon 12

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web video answer so now we ll talk about problems 58 from chapter four this problem whereas how do isotopes differ and how are they similar isotopes for the same element so isotopes differ in the number of neutro

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web isotopes are atoms of the same element that differ in the amount of neutrons and atomic mass almost all elements on the periodic table have at least two different natural isotopes many elements

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web the isotopes of a given element always contain the same number of protons and therefore occupy the same place on the periodic table however because isotopes contain different numbers of neutrons each isotope has a unique atomic mass

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web atoms of the same element i e same atomic number  $z$  that have different numbers of neutrons are called isotopes for example  $^{99}C$  of the carbon atoms on earth have 6 neutrons and 6 protons in their nuclei about 1 of the carbon atoms have 7

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web aug 19 2022 atoms with the same number of protons but different numbers of neutrons are called isotopes they share almost the same chemical properties but differ in mass and therefore in physical properties there are stable isotopes which do not emit radiation and there are unstable isotopes which do emit radiation

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web aug 9 2023 how do different isotopes of an element differ two isotopes differ in the number of neutrons in the nucleus for example carbon 14 has 8 neutrons differing from carbon 12 which has 6

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web atoms of the same element that differ in their numbers of neutrons are called isotopes many isotopes occur naturally usually one or two isotopes of an element are the most stable and common different isotopes of an element generally have the same physical and chemical properties because they have the same numbers of protons and electrons

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web jan 5 2016 explanation isotopes are atoms of the same element same atomic number ie same number of protons in the nucleus but different atomic masses different mass numbers due to different numbers of neutrons in the nucleus for example we may get an isotope of chlorine with 17 protons and 18 neutrons hence atomic mass 35

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nuclear particles i e z z 1 we gots hydrogen z 2 we gots helium z 3 lithium z 47 we gots silver the nucleus can also contain neutrons massive nuclear particles of zero charge

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web oct 31 2023 an isotope is one of two or more species of atoms of a chemical element with the same atomic number and position in the periodic table and nearly identical chemical behavior but with different atomic masses and physical properties every chemical element has one or more isotopes

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web jul 18 2018 answer link due to the number of neutrons isotopes differ with respect to the number of neutrons their nuclei contain where neutrons are massive particles of zero charge a nucleus may contain variable numbers of neutrons and this gives rise to the phenomenon of isotopes

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web part i answer the questions based on the above reading what is an isotope isotopes are versions of the same element they have the same number of protons and electrons as the element but different mass numbers and number of neutrons what does the number next to isotopes signify the number indicates the isotope's mass number

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web nov 27 2016 explanation basically isotopes are nuclei that have the same number of protons and are thus the same element but have unequal numbers of neutrons answer link the nucleus differs as a result of the number of neutrons in the atom the actual structure doesn't differ all that much but its mass does differ thus its density is a little

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