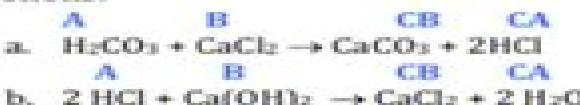


## Chapter 19 Study Guide: Acids, Bases & Titration – Answers

1. Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids and bases.

	<b>ACIDS</b>	<b>BASES</b>
<b>Arrhenius</b>	Form hydronium ions ( $\text{H}_3\text{O}^+$ ) in aqueous sol'n	Form hydroxide ions ( $\text{OH}^-$ ) in aqueous sol'n
<b>Brønsted-Lowry</b>	Proton ( $\text{H}^+$ ) donors	Proton ( $\text{H}^+$ ) acceptors
<b>Lewis</b>	Electron pair acceptors	Electron pair donors

2. Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:



3. Write the names or formulas for the following acids and bases:

a. $\text{H}_2\text{S}$ – <b>hydrosulfuric acid</b>	e. Nitrous acid – <b><math>\text{HNO}_2</math></b>
b. $\text{HC}_2\text{H}_3\text{O}_2$ – <b>acetic acid</b>	f. Calcium hydroxide – <b><math>\text{Ca}(\text{OH})_2</math></b>
c. $\text{NaOH}$ – <b>sodium hydroxide</b>	g. hydrochloric acid – <b><math>\text{HCl}</math></b>
d. $\text{HNO}_3$ – <b>nitric acid</b>	

Use the following formulas for the remaining questions: (these will not be given to you on the test!)

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$\text{M}_\text{A}\text{-Va}^- \text{M}_\text{B} = \text{M}_\text{B}\text{-V}_\text{a}^- \text{M}_\text{A}$$

4. What is the pH of  $1.34 \times 10^{-4}$  M solution of hydrochloric acid?

$$\text{pH} = -\log[1.34 \times 10^{-4}\text{M}] = 3.873$$

5. Determine the pH and pOH of a  $1.85 \times 10^{-6}$  M NaOH solution.

$$\text{pOH} = -\log[1.85 \times 10^{-6}\text{M}] = 5.733$$

$$\text{pH} = 14 - 5.73 = 8.267$$

6. What is the  $[\text{OH}^-]$  of an aqueous solution that has a pH of 5.0?

$$\text{pOH} = 9 \quad [\text{OH}^-] = 1 \times 10^{-9}\text{M}$$

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35\text{gNaOH}}{2.000\text{L}} \times \frac{1\text{molNaOH}}{40.00\text{gNaOH}} = 0.44\text{M NaOH}$$

$$\text{pOH} = -\log[0.44\text{M}] = 0.36 \quad \text{pH} = 14 - 0.36 = 13.64 \quad \text{basic}$$

# Acids And Bases Exam Review Guide Answers

**Ralph H. Petrucci**

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## Table of Contents Acids And Bases Exam Review Guide Answers

1. Understanding the eBook Acids And Bases Exam Review Guide Answers
  - The Rise of Digital Reading Acids And Bases Exam Review Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Acids And Bases Exam Review Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Acids And Bases Exam Review Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Acids And Bases Exam Review Guide Answers
  - Personalized Recommendations
  - Acids And Bases Exam Review Guide Answers User Reviews and Ratings
  - Acids And Bases Exam Review Guide Answers and Bestseller Lists
5. Accessing Acids And Bases Exam Review Guide Answers Free and Paid eBooks
  - Acids And Bases Exam Review Guide Answers Public Domain eBooks
  - Acids And Bases Exam Review Guide Answers eBook Subscription Services
  - Acids And Bases Exam Review Guide Answers Budget-Friendly Options
6. Navigating Acids And Bases Exam Review Guide Answers eBook Formats
  - ePUB, PDF, MOBI, and More
  - Acids And Bases Exam Review Guide Answers Compatibility with Devices
  - Acids And Bases Exam Review Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Acids And Bases Exam Review Guide Answers
  - Highlighting and Note-Taking Acids And Bases Exam Review Guide Answers
  - Interactive Elements Acids And Bases Exam Review Guide Answers
8. Staying Engaged with Acids And Bases Exam Review Guide Answers

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Acids And Bases Exam Review Guide Answers

9. Balancing eBooks and Physical Books Acids And Bases Exam Review Guide Answers

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Acids And Bases Exam Review Guide Answers

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Acids And Bases Exam Review Guide Answers

- Setting Reading Goals Acids And Bases Exam Review Guide Answers
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Acids And Bases Exam Review Guide Answers

- Fact-Checking eBook Content of Acids And Bases Exam Review Guide Answers
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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