

Ch10 Study Guide

Names

Period

Chemists relate units of counting, of mass, and of volume to a single quantity called the 1. The number of representative particles in one mole of a substance is 2. To find the mass of a mole of a compound, scientists add together the 3 of the atoms in the compound. When you substitute the unit "grams" for amu, you obtain the 4 of the compound. There are 5 representative particles in a mole of any substance.

1	2	3	4	5
mole	Avagadro number	atomic mass	molar mass	6.02×10^{23}

AI-NT-ST

ST 6. a mole of a pure substance contains 6.02×10^{23} atoms.

ST 7. The representative particle of a compound is the molecule.

AT 8 A mole of CCl_4 is composed of one atom of carbon and 4 atoms of chlorine.

AT 9. A mole of carbon atoms has a mass approximately 3 times as great as the mass of a mole of helium atoms.

NT 10. The molar mass of nitrogen gas, N_2 , is 14.02 g.

March

B 11. Avogadro's number a) atoms, molecules, or ions in a substance

C 12. molar mass b) 6.02×10^{23}

D 13. mole c) the mass of 1 mole of a substance

14. representative particles d) SI unit that measures the amount of a substance

Find the molar mass of a) NaNO_3 b) $\text{Cr}_2(\text{SO}_4)_3$ c) $(\text{NH}_4)_2\text{CO}_3$ d) N_2O_5

$$A=22.99+14.01+16.00+16.00+16.00=85 \text{ g/mol}$$
$$B: 52.00 + 52.00 + 52.00 + 32.06 + 32.06 + 16.00 + 16.00 + 16.00 + 16.00 + 16.00 + 16.00 + 16.00 + 16.00 = 348.12 \text{ g/mol}$$
[illegible]

a) 85g/mol b) 348.12g/mol c) 96.11g/mol d) 242.07g/mol

D-
14.01+14.01+14.01+14.01+14.01+14.01+14.01+16.00+16.00+16.00+16.00+16.00+16.00+16.00+16.00+16.00=242.07g
mol

Bju Chemistry Study Guide

Steven S. Zumdahl, Paul B. Kelter



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Catalogs of the Scripps Institution of Oceanography Library Scripps Institution of Oceanography. Library, 1980

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