

# Basic Chemistry Study Guide 2

## 1. Hydrogen bonds

1. Form between two polar molecules **covalent bonds**
2. A hydrogen's &+ region is attracted to a &- region in another molecule.
3. Form between water molecules (Where are the &+ and &- regions in a water molecule? **hydrogen bonds, at the top of the symbol**)
4. Each hydrogen bond is very weak (can be easily formed and broken).
5. There are many hydrogen bonds in water. Together, these bonds give water its unique properties that are essential for life.
6. One of these properties is the ability to be a solvent for other polar molecules.
7. Hydrogen bonds are depicted in figures as dotted or dashed lines between two atoms.

## 2. Solutions, solvents, and solutes

1. Water is the most abundant molecule in living organisms, and 60% of human body weight is water.
2. Water is located inside and outside of our cells.
3. Body fluids are solutions.
4. Define the terms **solution**, **solvent**, and **solute** **Solvent: a substance capable of dissolving other polar molecules and ionic bonds; solutes: molecules dissolved in the solvent, solution: homogeneous mixture**

## 3. How substances dissolve in water

1. When a solute dissolves, it becomes surrounded by water molecules.
2. For a solute to dissolve, it must attract water molecules.
3. What kinds of substances attract water molecules? **Sugars, nucleic acids, and amino acids**
4. When a solute dissolves it forms hydrogen bonds with water molecules.
5. Define the terms **ion**, **cation**, and **anion**. **ion: atom with a charge; cation: positively charged ion, anions: negatively charged ion**
6. Define the terms **hydrophilic** and **hydrophobic** **Hydrophilic: attracted to water; hydrophobic: afraid of water**
7. What types of molecules are hydrophilic and what types are hydrophobic? **Nonpolar: hydrophobic, polar: hydrophilic**

# Basic Chemistry Study Guide For Anatomy

**DP Hallahan**

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## **Table of Contents Basic Chemistry Study Guide For Anatomy**

1. Understanding the eBook Basic Chemistry Study Guide For Anatomy
  - The Rise of Digital Reading Basic Chemistry Study Guide For Anatomy
  - Advantages of eBooks Over Traditional Books
2. Identifying Basic Chemistry Study Guide For Anatomy
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Basic Chemistry Study Guide For Anatomy
  - User-Friendly Interface
4. Exploring eBook Recommendations from Basic Chemistry Study Guide For Anatomy
  - Personalized Recommendations
  - Basic Chemistry Study Guide For Anatomy User Reviews and Ratings
  - Basic Chemistry Study Guide For Anatomy and Bestseller Lists
5. Accessing Basic Chemistry Study Guide For Anatomy Free and Paid eBooks
  - Basic Chemistry Study Guide For Anatomy Public Domain eBooks
  - Basic Chemistry Study Guide For Anatomy eBook Subscription Services
  - Basic Chemistry Study Guide For Anatomy Budget-Friendly Options
6. Navigating Basic Chemistry Study Guide For Anatomy eBook Formats

- ePub, PDF, MOBI, and More
- Basic Chemistry Study Guide For Anatomy Compatibility with Devices
- Basic Chemistry Study Guide For Anatomy Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Basic Chemistry Study Guide For Anatomy
- Highlighting and Note-Taking Basic Chemistry Study Guide For Anatomy
- Interactive Elements Basic Chemistry Study Guide For Anatomy

8. Staying Engaged with Basic Chemistry Study Guide For Anatomy

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Basic Chemistry Study Guide For Anatomy

9. Balancing eBooks and Physical Books Basic Chemistry Study Guide For Anatomy

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Basic Chemistry Study Guide For Anatomy

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Basic Chemistry Study Guide For Anatomy

- Setting Reading Goals Basic Chemistry Study Guide For Anatomy
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Basic Chemistry Study Guide For Anatomy

- Fact-Checking eBook Content of Basic Chemistry Study Guide For Anatomy
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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