

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and K_a or K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|-------------------|-----------------------------|---------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. neutral solution |
| | 7. acid solution | 8. pH |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) for the reaction(s) indicated. Show the calculation for pH or percent dissociation where indicated. Assume a 0.10 M concentration unless noted.

1. What substance is H^+ in water? Hydrogen ion is H_3O^+ .
2. Write the equation for the autoionization of water which produces an H^+ ion and a hydroxide ion. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$
3. Hydrochloric acid dissociates from HCl to H^+ and Cl^- . Write the K_a expression for this reaction. $K_a = \frac{[\text{H}^+][\text{Cl}^-]}{[\text{HCl}]}$
4. Write the K_b expression for the reaction of NH_3 with H_2O . $K_b = \frac{[\text{NH}_4^+][\text{OH}^-]}{[\text{NH}_3]}$
5. Write the K_a expression for the reaction of H_2O with H_2O . $K_a = \frac{[\text{H}^+][\text{OH}^-]}{[\text{H}_2\text{O}]}$
6. Write the K_b expression for the reaction of H_2O with H_2O . $K_b = \frac{[\text{H}^+][\text{OH}^-]}{[\text{H}_2\text{O}]}$

7. HCl is a strong acid.
8. HCl is a strong acid.
9. HCl is a strong acid.
10. HCl is a strong acid.
11. HCl is a strong acid.
12. HCl is a strong acid.
13. HCl is a strong acid.
14. HCl is a strong acid.
15. HCl is a strong acid.
16. HCl is a strong acid.
17. HCl is a strong acid.
18. HCl is a strong acid.

19 2 Hydrogen Ions And Acidity Work Sheet Answers

Rod O'Connor



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