

Example 15.1 Identifying Brønsted–Lowry Acids and Bases and Their Conjugates

In each reaction, identify the Brønsted–Lowry acid, the Brønsted–Lowry base, the conjugate acid, and the conjugate base.



Solution

a. Because H_2SO_4 donates a proton to H_2O in this reaction, it is the acid (proton donor). After H_2SO_4 donates the proton, it becomes HSO_4^- , the conjugate base. Because H_2O accepts a proton, it is the base (proton acceptor). After H_2O accepts the proton, it becomes H_3O^+ , the conjugate acid.



Acid	Base	Conjugate base	Conjugate acid
$\text{H}_2\text{SO}_4(aq)$	$\text{H}_2\text{O}(l)$	$\text{HSO}_4^-(aq)$	$\text{H}_3\text{O}^+(aq)$

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