



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Explain how the concentration of hydrogen ions in a solution is measured.
- Describe strong and weak acids.

Key Vocabulary

<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base	<ul style="list-style-type: none">concentrationhydrogen ionionlogarithmlogarithmic scaleacidstrong acidweak acidbasestrong baseweak basepHhydroxide ionconjugate acidconjugate base
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Key Equations

- $\text{pH} = -\log[\text{H}^+]$
- $\text{pH} = \log \frac{1}{[\text{H}^+]}$

Key Figures

- Figure 1: pH scale
- Figure 2: Ionization of water

Principles of Ionization

Acid dissociation is the process of ionization in which an acid dissociates into ions. This process is reversible.

Water dissociation is the process of ionization in which water dissociates into ions. This process is reversible.

Strong acids dissociate completely in water. Strong acids are strong electrolytes.

Weak acids dissociate partially in water. Weak acids are weak electrolytes.

Strong Acids

- Hydrochloric acid
- Sulfuric acid
- Perchloric acid
- Hydrobromic acid
- Hydroiodic acid
- Hydrofluoric acid
- Acetic acid
- Chloroacetic acid
- Perchloroacetic acid
- Hydrocyanic acid
- Hydrochlorous acid
- Hydrobromous acid
- Hydroiodous acid

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