

Chemical Reactions:

1. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. What is the ΔH°_f for the reaction?
- A) greater than 2.00 kcal/mol
 B) greater than 1.00 kcal/mol
 C) less than 1.00 kcal/mol
 D) less than 2.00 kcal/mol
 E) greater than 4.00 kcal/mol
2. A 1.0-mol sample of PCl_3 reacts with excess oxygen according to the equation: $2\text{PCl}_3(\text{g}) \rightarrow 2\text{PCl}_5(\text{g}) + \text{O}_2(\text{g})$. Below, which reaction will yield the greatest amount of O_2 after PCl_3 is removed?
- A) 0.0 mol
 B) 1.0 mol
 C) 1.2 mol
 D) 1.7 mol
 E) 2.0 mol
3. The reaction of NO_2 is step characterized by the equation: $4\text{NO}_2 \rightarrow 3\text{O}_2 + 2\text{N}_2\text{O}_5$. If 1.0 mol of NO_2 is added to 1.0 mol of N_2O_5 and the reaction is allowed to proceed, how will the reaction rate change?
- A) 1.000 mol/l
 B) 0.750 mol/l
 C) 1.000 mol/l
 D) 1.50 mol/l
 E) 2.00 mol/l
4. Calculated are the following reactions, or from the experiments, are unknowns, calculated.
- $\text{H}_2\text{O}_2(\text{l}) + \text{H}_2\text{O}_2(\text{l}) \rightarrow 2\text{H}_2\text{O}_2(\text{l}) = 300\text{J/g}$
- $\text{H}_2 + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$ (enthalpy of difference, ΔH°_f for production of 1.00 mol of H_2O_2 , what is the ΔH°_f for H_2 ?)
- A) 0.0
 B) 2.00
 C) 0
 D) 3.0
 E) 0.5
5. Refer to the following reaction:
- $\text{NO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightarrow 4\text{NO}(\text{g}) + \text{O}_2(\text{g})$
- Major steady-state levels of NO_2 are given below. In which below is ΔH°_f most likely to be closest to ΔH°_f for the reaction?
- A) 1.500
 B) 1.750×10^3
 C) 1.800×10^3
 D) 1.700
 E) Insufficient information

6. The following reaction $2\text{H}_2\text{O}_2 \rightarrow 2\text{H}_2\text{O} + \text{O}_2$ is given. Below, which reaction will yield the greatest amount of O_2 after 1.00 mol of hydrogen and oxygen react?

- A) 1.000 mol/l
 B) 1.200 mol/l
 C) 1.750 mol/l
 D) 2.000 mol/l
 E) 2.500 mol/l

7. Refer to the following reaction:
- $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{OH} + \text{H}_2\text{O}$
- What mass of oxygen, O_2 , is required to react completely with 1.00 g of C_2H_6 ?
- A) 1.000×10^{-3} g
 B) 1.000×10^{-2} g
 C) 1.000×10^{-1} g
 D) 1.000×10^0 g
 E) 1.000×10^1 g

8. After reactor initialization and shutdown of reactor, the air is removed at 770°C and prepared to become the following reaction:



- A) 1.000×10^{12} mol/electron
 B) 1.000×10^{13} mol/electron
 C) 1.000×10^{14} mol/electron
 D) 1.000×10^{15} mol/electron
 E) 1.000×10^{16} mol/electron

9. Calculate the mass of carbon dioxide produced from 1.00 g of octane, C_8H_{18} , in the following reaction:



- A) 1.75×10^{-3} g
 B) 1.75×10^{-2} g
 C) 1.75×10^{-1} g
 D) 1.75×10^0 g
 E) 1.75×10^1 g

10. Calculate the mass of oxygen required to react with 1.00 g of sulfur in the following reaction:



- A) 1.000×10^{12} mol/electron O_2
 B) 1.000×10^{13} mol/electron O_2
 C) 1.000×10^{14} mol/electron O_2
 D) 1.000×10^{15} mol/electron O_2
 E) 1.000×10^{16} mol/electron O_2

Multiple Choice Stoichiometry Test

**James S. Mattson, Harry B.
Mark, Hubert C. MacDonald**

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Table of Contents Multiple Choice Stoichiometry Test

1. Understanding the eBook Multiple Choice Stoichiometry Test
 - The Rise of Digital Reading Multiple Choice Stoichiometry Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Stoichiometry Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Multiple Choice Stoichiometry Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Stoichiometry Test
 - Personalized Recommendations
 - Multiple Choice Stoichiometry Test User Reviews and Ratings
 - Multiple Choice Stoichiometry Test and Bestseller Lists
5. Accessing Multiple Choice Stoichiometry Test Free and Paid eBooks
 - Multiple Choice Stoichiometry Test Public Domain eBooks
 - Multiple Choice Stoichiometry Test eBook Subscription Services
 - Multiple Choice Stoichiometry Test Budget-Friendly Options

6. Navigating Multiple Choice Stoichiometry Test eBook Formats
 - ePUB, PDF, MOBI, and More
 - Multiple Choice Stoichiometry Test Compatibility with Devices
 - Multiple Choice Stoichiometry Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Multiple Choice Stoichiometry Test
 - Highlighting and Note-Taking Multiple Choice Stoichiometry Test
 - Interactive Elements Multiple Choice Stoichiometry Test
8. Staying Engaged with Multiple Choice Stoichiometry Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Multiple Choice Stoichiometry Test
9. Balancing eBooks and Physical Books Multiple Choice Stoichiometry Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Multiple Choice Stoichiometry Test
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Stoichiometry Test
 - Setting Reading Goals Multiple Choice Stoichiometry Test
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Stoichiometry Test
 - Fact-Checking eBook Content of Multiple Choice Stoichiometry Test
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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