

Chemistry 2nd Semester Exam Review Sheet:**Chapters 9, 10, 11, 12, 13, 14, 15, and 16****Chapter 9: Chemical Quantities**

1. Solve stoichiometry problems using grams \rightarrow moles \rightarrow moles \rightarrow grams calculations. Examples below:
a. Calculate moles/grams of $(\text{NH}_4)_2\text{SO}_4$ produced when 30.0 moles of NH_3 react with H_2SO_4 .



$$30.0 \text{ moles} \times 1 \text{ mol/2 mol} = 15.0 \text{ moles}$$

$$15.0 \text{ moles } (\text{NH}_4)_2\text{SO}_4 \times 132.2 \text{ g/mol/mole} = 1983 = 1980 \text{ grams (to 3 s.f.)}$$

b. $\text{Na}_2\text{CO}_3 + 2\text{HNO}_3 \rightarrow 2\text{NaNO}_3 + \text{CO}_2 + \text{H}_2\text{O}$. How many moles/grams of sodium nitrate are produced by the complete reaction of 100. grams of sodium carbonate? Moles of nitric acid needed to produce 500.0 grams of sodium nitrate? *Be able to calculate grams or moles of any reactant or product given a quantity of reactant or product.*

Grams \rightarrow Moles \rightarrow Moles \rightarrow Grams

$$100. \text{ g } \text{Na}_2\text{CO}_3 \times 1 \text{ mol/106.0 g} \times 2 \text{ mol } \text{NaNO}_3 / 1 \text{ mol } \text{Na}_2\text{CO}_3 \times 85.00 \text{ g/mol} = 160.4 \text{ g}$$

$$500.0 \text{ g } \text{NaNO}_3 \times 1 \text{ mol/85.00 g} \times 2 \text{ mol / 2 mol} \times 63.02 \text{ g/mol} = 370.7 \text{ grams}$$

2. Calculate percent yield of products. Example below using (b) above.

What is the theoretical yield of NaNO_3 if 100.0 grams of Na_2CO_3 are used in the reaction? What is the percent yield of NaNO_3 if 130.5 grams of NaNO_3 are actually collected?

Theoretical yield = 160.4 grams (answer from above)

$$\text{Percent Yield} = (\text{Actual Yield/Theoretical Yield}) \times 100 = (130.5 \text{ g}/160.4 \text{ g}) \times 100 = 81.36\%$$

3. What is a limiting reactant in a chemical reaction? The reactant that is used up first, thereby limiting the amount of product that can be formed.

Chapter 10: Modern Atomic Theory

4. Describe Rutherford's model of the atom. What could it not explain?

Nucleus (very small) consisting of protons with electrons in orbits around the nucleus. Could not explain how the electrons remained in orbit.

5. Define electromagnetic radiation, wavelength, and frequency.

6. How do atoms gain energy to become excited? Absorb a photon of energy.

How do they lose energy to return to their ground state? Release a photon of same energy as that absorbed.

7. What is the evidence that energy absorbed by the hydrogen atom is quantized? Specific wavelengths of light seen in emission spectra for atoms (page 285-86)

8. Describe the Bohr model of the atom. How did it improve on Rutherford's model?

Same as Rutherford with respect to nucleus. Electrons located on orbits at specific distances from nucleus.

9. Draw the orbital diagram (with the arrows) for an element: for example, phosphorus, sodium, and fluorine.

See book, notes, probs 54, 96

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Daniela Niemeyer

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Dehn,2011-09-07 A hands on memory training program for children and adolescents featuring dozens of practical evidence based memory exercises A practical workbook designed to assist students whose academic learning is suffering due to a memory deficit or ineffective utilization of their memory capabilities Helping Students Remember provides numerous strategies and methods to strengthen memory including chunking organization keyword self testing pegword loci and mnemonics Drawing on the author s extensive training and experience this useful resource presents effective techniques and lessons on How memory works Memorization methods Goals for improving memory Repetition Using cards to build memory Grouping words by category Study skills that help memory Using arithmetic to build memory Using music to remember Improving recall during tests Creating and using review sheets Picturing verbal information Using context cues Plans for using memory strategies With an accompanying CD containing all of the worksheets and word lists for reproduction Helping Students Remember is the first workbook of its kind for general psychologists school psychologists and special education teachers offering practical easy to implement and evidence based methods for working with children with memory impairments

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1. Understanding the eBook 2nd Semester Exam Review Sheet Answer
 - The Rise of Digital Reading 2nd Semester Exam Review Sheet Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying 2nd Semester Exam Review Sheet Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 2nd Semester Exam Review Sheet Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from 2nd Semester Exam Review Sheet Answer

- Personalized Recommendations
- 2nd Semester Exam Review Sheet Answer User Reviews and Ratings
- 2nd Semester Exam Review Sheet Answer and Bestseller Lists

5. Accessing 2nd Semester Exam Review Sheet Answer Free and Paid eBooks

- 2nd Semester Exam Review Sheet Answer Public Domain eBooks
- 2nd Semester Exam Review Sheet Answer eBook Subscription Services
- 2nd Semester Exam Review Sheet Answer Budget-Friendly Options

6. Navigating 2nd Semester Exam Review Sheet Answer eBook Formats

- ePub, PDF, MOBI, and More
- 2nd Semester Exam Review Sheet Answer Compatibility with Devices
- 2nd Semester Exam Review Sheet Answer Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 2nd Semester Exam Review Sheet Answer
- Highlighting and Note-Taking 2nd Semester Exam Review Sheet Answer
- Interactive Elements 2nd Semester Exam Review Sheet Answer

8. Staying Engaged with 2nd Semester Exam Review Sheet Answer

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers 2nd Semester Exam Review Sheet Answer

9. Balancing eBooks and Physical Books 2nd Semester Exam Review Sheet Answer

- Benefits of a Digital Library
- Creating a Diverse Reading Collection 2nd Semester Exam Review Sheet Answer

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine 2nd Semester Exam Review Sheet Answer

- Setting Reading Goals 2nd Semester Exam Review Sheet Answer
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of 2nd Semester Exam Review Sheet Answer

- Fact-Checking eBook Content of 2nd Semester Exam Review Sheet Answer
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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