

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on the pH and  $K_a$  or  $K_b$  and vice versa.
3. Calculate the pH of weak acids and bases.

## Vocabulary

- |                   |                             |                  |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. acid rain                | 6. acid rain     |
| 7. autoionization | 8. acid rain                | 9. acid          |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the equation for the autoionization of water. What is the  $K_w$  for water at 25°C? What is the  $\text{pH}$  of pure water at 25°C? What is the  $\text{pH}$  of pure water at 100°C?

1. What is the  $\text{pH}$  of a 0.1 M solution of hydrochloric acid?
2. What is the  $\text{pH}$  of a 0.1 M solution of sodium hydroxide?
3. What is the  $\text{pH}$  of a 0.1 M solution of acetic acid?
4. What is the  $\text{pH}$  of a 0.1 M solution of ammonia?
5. What is the  $\text{pH}$  of a 0.1 M solution of hydrofluoric acid?
6. What is the  $\text{pH}$  of a 0.1 M solution of hydrocyanic acid?
7. What is the  $\text{pH}$  of a 0.1 M solution of hydrobromic acid?
8. What is the  $\text{pH}$  of a 0.1 M solution of hydroiodic acid?
9. What is the  $\text{pH}$  of a 0.1 M solution of hydrochloric acid?
10. What is the  $\text{pH}$  of a 0.1 M solution of hydrobromic acid?
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12. What is the  $\text{pH}$  of a 0.1 M solution of hydrobromic acid?
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# 19 2 Hydrogen Ions And Acidity Answer Key

**Sebastian Brünink**



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