

14.1

HYDROGEN IONS AND ACIDITY

Section Overview

Objectives

1. Identify conjugate acid-base pairs of ions, acids, or hydroxides and of species associated with them.
2. Calculate hydroxide ion concentrations, hydronium ion concentrations, and pH values for weak acids and bases.
3. Calculate the pH of weak acid solutions.

Vocabulary

- | | | |
|-------------------|-----------------------------|------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. acid solution |
| 4. autoionization | 5. hydronium ion | 6. hydroxide ion |
| | 7. weak acid | 8. pH |

Key Equations

- | | |
|--|---|
| 1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ (14.1) | 2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) |
| 3. $\text{pH} = -\log[\text{H}^+]$ (14.3) | 4. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ (14.4) |

Part A: Completion

Use the appropriate number to identify the relationship of the species indicated that are associated in the reaction. Also indicate which species acts as the conjugate acid.

1. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) (14.1) (14.3) (14.4)
2. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) (14.1) (14.3) (14.4)
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10. $\text{H}_2\text{O} + \text{H}_2\text{O} \rightleftharpoons \text{H}_3\text{O}^+ + \text{OH}^-$ (14.2) (14.1) (14.3) (14.4)

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19 2 Hydrogen Ions And Acidity Worksheet Answers

W. M. Clark



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