

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and pOH , and vice versa from pH and pOH .
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pH | 6. pOH |

Key Equation

- | | |
|-------------------------------------|----------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = 14 - \text{pOH}$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = 14 - \text{pOH}$ |

Part A Completion

Write the hydrogen ion concentration or autoionization equilibrium of the acids and bases that are indicated in the spaces provided. Be sure to include the correct state symbols.

1. What substance is _____, a weak hydrogen ion concentration of 10^{-4} M? The autoionization of this acid is _____.
2. Hydrochloric acid has a pH of _____, or a pOH of _____.
3. _____, a concentrated solution of the base, has a pH of _____.
4. _____, a weak acid, has a pH of _____.
5. _____, a weak base, has a pOH of _____.
6. _____, a weak acid, has a pH of _____.
7. _____, a weak base, has a pOH of _____.
8. _____, a weak acid, has a pH of _____.

9. _____
10. _____
11. _____
12. _____
13. _____
14. _____
15. _____
16. _____
17. _____
18. _____

19 2 Hydrogen Ions And Acidity Work Sheet Answers

Horace Willard Davenport



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