

Introduction:

When we bake/cook something, we use a specific amount of each ingredient. Imagine if you made a batch of cookies and used way too many eggs, or not enough sugar. YUCK! In chemistry, reactions proceed with very specific recipes. The study of these recipes is stoichiometry. When the reactants are present in the correct amounts, the reaction makes products. We are going to look at what happens when there are more or less of some of the reactants present.

Procedure: PHET Simulation: [Play with the Sims of Chemistry](#) → [Acid-base](#), [Predictors](#), and [Leftovers](#) OPEN IN HTML5

[https://phet.colorado.edu/en/simulations/versions/predictors-products-and-leftovers?sim=reactants-products-and-leftovers_molecules](#)

Part 1: Making Sandwiches



1. A cheese sandwich requires 2 slices of bread to 1 slice of cheese. If there is exactly twice as many slices of bread as cheese, then there will be no leftovers. What happens if you have two slices of bread and one slice of cheese? You have a slice of cheese left over!
2. Use the simulation to play around with this two-reactant synthesis reaction. Take some time and familiarize yourself with the simulation buttons.
3. At the bottom of the screen, change the coefficients of the reactants to 2:1.



4. Notice that there is now one sandwich with no leftovers!
5. Complete situations B through D. A has been done for you as an example.

Situation	Bread Used	Cheese Used	Sandwiches Made	Excess Bread	Excess Cheese
A	4 slices	2 slices	2 sandwiches	0 bread	0 cheese
B	2 slices	3 slices			
C			2 sandwiches	1 slice	0 slices
D	6 slices		3 sandwiches		4 slices

6. Look at situation A. The cheese in this case is said to be in excess. The other reactant (bread) has no leftovers. The bread in this case is the *limiting* reactant.
7. In situations B, C, and D, determine the limiting reactant: B: _____ C: _____ D: _____



Part 2: Real Chemical Reactions



8. Select "Molecules" towards the bottom of your screen. We can now work with real chemical reactions, one that creates a very contaminating BCl3!
9. What is the lowest molecular ratio for the reaction of hydrogen and oxygen to produce water? That is, balance the following equation:



10. Complete the table below while making water H2O from hydrogen H2 and oxygen O2.

Hydrogen Molecules <chem>H2</chem>	Oxygen Molecules <chem>O2</chem>	Water Molecules <chem>H2O</chem>	Excess <chem>H2</chem>	Excess <chem>O2</chem>
4 molecules	4 molecules			
7 molecules	6 molecules			
		4 molecules	0 molecules	0 molecules
6 molecules	8 molecules			
		4 molecules	1 molecules	0 molecules

Stop for check point. Teacher signature for this page: _____

Basic Stoichiometry Phet Lab Answer Key

JG Myers



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