

BIOCHEMISTRY

CHEMICAL CONCEPTS



PERIODIC TABLE

Groups: 1, 2, 13, 14, 15, 16, 17, 18

Periods: 1, 2, 3, 4, 5, 6, 7

Metals, Non-metals, Metalloids



BIOCHEMICAL PRINCIPLES

1. Carbohydrates

Carbohydrates are organic compounds consisting of carbon, hydrogen, and oxygen. They are classified into monosaccharides, disaccharides, and polysaccharides.

Monosaccharides: Simple sugars that cannot be hydrolyzed into simpler sugars. Examples include glucose, fructose, and galactose.

Disaccharides: Composed of two monosaccharide units linked by a glycosidic bond. Examples include sucrose, lactose, and maltose.

Polysaccharides: Long chains of monosaccharide units. Examples include starch, glycogen, and cellulose.

2. Lipids

Lipids are hydrophobic or amphipathic molecules. They are classified into fats, oils, and steroids.

Fats and Oils: Composed of a glycerol backbone and fatty acid chains. Saturated fats are solid at room temperature, while unsaturated fats are liquid.

Steroids: Have a characteristic four-ring structure. Examples include cholesterol and sex hormones.

3. Proteins

Proteins are polymers of amino acids. They are essential for the structure and function of all living organisms.

Amino Acids: The building blocks of proteins. Each amino acid has a central carbon atom bonded to a hydrogen atom, an amino group, a carboxyl group, and a variable R-group.

Protein Structure: Proteins can have primary, secondary, tertiary, and quaternary structures.

4. Enzymes

Enzymes are biological catalysts that speed up chemical reactions. They are proteins with specific active sites.

Enzyme Activity: Measured by the rate of reaction. Factors affecting enzyme activity include temperature, pH, and substrate concentration.

Enzyme Inhibition: Can be reversible or irreversible. Competitive inhibitors compete with the substrate for the active site.

5. Vitamins

Vitamins are organic compounds that are essential for normal growth and metabolism. They are classified into fat-soluble (A, D, E, K) and water-soluble (B, C) vitamins.

6. Types of Chemical Groups

Hydroxyl Group (-OH): Found in alcohols and carbohydrates. It is polar and can form hydrogen bonds.

Carboxyl Group (-COOH): Found in amino acids and organic acids. It can donate protons and form hydrogen bonds.

Amino Group (-NH2): Found in amino acids. It can accept protons and form hydrogen bonds.

Phosphate Group (-PO4): Found in nucleic acids and energy carriers like ATP.

7. Nucleic Acids

Nucleic acids are polymers of nucleotides. They store and transmit genetic information.

Deoxyribonucleic Acid (DNA): The genetic material in most organisms. It is a double helix structure.

Ribonucleic Acid (RNA): Involved in protein synthesis. It is a single helix structure.

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8. Water

Water is a polar molecule and is essential for life. It acts as a solvent and is involved in many biological processes.

Hydrogen Bonding: Water molecules are held together by hydrogen bonds, which are responsible for its unique properties like high boiling point and surface tension.

9. Buffers

Buffers are solutions that resist changes in pH. They are important for maintaining the pH of biological systems.

Acid-Base Balance: The body maintains a constant pH through the action of buffers and the respiratory and renal systems.

10. Osmosis and Diffusion

Osmosis is the movement of water across a semi-permeable membrane. Diffusion is the movement of solutes from high to low concentration.

Cellular Transport: Cells use osmosis and diffusion to maintain their internal environment.

11. Energy

Energy is required for all biological processes. It is stored in molecules like ATP and released during metabolism.

Metabolism: The sum of all chemical reactions in the body. It includes catabolism (breaking down) and anabolism (building up).

12. Homeostasis

Homeostasis is the maintenance of a stable internal environment. It involves the regulation of temperature, pH, and fluid balance.

Feedback Mechanisms: The body uses negative feedback loops to maintain homeostasis.

13. Genetic Inheritance

Genetic inheritance is the passing of traits from parents to offspring. It is controlled by genes on chromosomes.

Mendel's Laws: The basic principles of inheritance discovered by Gregor Mendel.

14. Evolution

Evolution is the change in the characteristics of a population over time. It is driven by natural selection and genetic drift.

Speciation: The process by which new species are formed.

15. Biotechnology

Biotechnology is the use of living organisms to produce products. It includes genetic engineering and fermentation.

Applications: Biotechnology is used in medicine, agriculture, and industry.

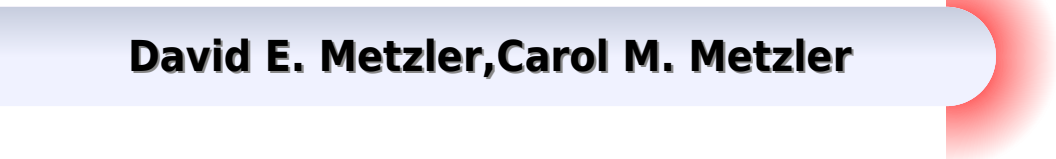
16. Environmental Science

Environmental science is the study of the interactions between the physical and biological environments. It focuses on pollution, climate change, and conservation.

Sustainability: Meeting the needs of the present without compromising the ability of future generations to meet their own needs.

Biochemistry Quickstudy Academic

David E. Metzler, Carol M. Metzler



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