

Chemistry Study Guide

Periodic Table of the Elements

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<small>3</small> Li <small>6.941</small>	<small>4</small> Be <small>9.012</small>																	<small>5</small> B <small>10.81</small>	<small>6</small> C <small>12.01</small>	<small>7</small> N <small>14.01</small>	<small>8</small> O <small>16.00</small>	<small>9</small> F <small>18.998</small>	<small>10</small> Ne <small>20.18</small>																									
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<small>19</small> K <small>39.10</small>	<small>20</small> Ca <small>40.08</small>	<small>21</small> Sc <small>44.96</small>	<small>22</small> Ti <small>47.88</small>	<small>23</small> V <small>50.94</small>	<small>24</small> Cr <small>52.00</small>	<small>25</small> Mn <small>54.94</small>	<small>26</small> Fe <small>55.85</small>	<small>27</small> Co <small>58.93</small>	<small>28</small> Ni <small>58.71</small>	<small>29</small> Cu <small>63.55</small>	<small>30</small> Zn <small>65.38</small>	<small>31</small> Ga <small>69.72</small>	<small>32</small> Ge <small>72.64</small>	<small>33</small> As <small>74.92</small>	<small>34</small> Se <small>78.96</small>	<small>35</small> Br <small>79.90</small>	<small>36</small> Kr <small>83.80</small>																															
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Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles 🎯

Matter:

Anything that has **mass** and **occupies space** ✖️🌐

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** 🧪👉

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

Section 2 nuclear Chemistry Study Guide Key

Charles E. Mortimer



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