

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids and bases.

Vocabulary

- | | | |
|------------------------|-----------------------------|-------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. ionization | 5. K_a | 6. K_b |
| 7. ionization constant | 8. pH | |

Key Equation

- | | |
|-----------------------|-----------------------|
| 1. $[H^+] = 10^{-pH}$ | 2. $pH = -\log [H^+]$ |
| 3. $pH = -\log [H^+]$ | 4. $[H^+] = 10^{-pH}$ |

Part A Completion

Write the equation and/or the numerical value(s) of the K_a and K_b for the acids and bases listed in the table. Also indicate which ionization constant is the stronger one.

1. HNO_3 is a strong acid. K_a is very large.
2. HCl is a strong acid. K_a is very large.
3. H_2SO_4 is a strong acid. K_a is very large.
4. H_2SO_3 is a weak acid. K_a is small.
5. H_2CO_3 is a weak acid. K_a is small.
6. H_2O is a weak acid. K_a is small.
7. H_2O is a weak base. K_b is small.
8. H_2O is a weak acid. K_a is small.
9. H_2O is a weak base. K_b is small.
10. H_2O is a weak acid. K_a is small.
11. H_2O is a weak base. K_b is small.
12. H_2O is a weak acid. K_a is small.
13. H_2O is a weak base. K_b is small.
14. H_2O is a weak acid. K_a is small.
15. H_2O is a weak base. K_b is small.

1.	HNO_3
2.	HCl
3.	H_2SO_4
4.	H_2SO_3
5.	H_2CO_3
6.	H_2O
7.	H_2O
8.	H_2O
9.	H_2O
10.	H_2O
11.	H_2O
12.	H_2O
13.	H_2O
14.	H_2O
15.	H_2O

19 2 Hydrogen Ions And Acidity Work Sheet Answers

W. M. Clark



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