



HYDROGEN IONS AND ACIDITY

Section Summary

Objectives

- Define hydrogen ion concentration and relate it to the pH of a solution.
- Define acid and base and relate them to the pH of a solution.
- Define strong and weak acids.

Key Vocabulary

| | |
|---|---|
| <ul style="list-style-type: none">hydrogen ionacidbasestrong acidweak acidpH | <ul style="list-style-type: none">hydrogen ion concentrationacidicbasicneutraltitrationtitrant |
|---|---|

Key Equations

- $\text{pH} = -\log_{10}[\text{H}^+]$
- $\text{pH} = \log_{10}[\text{OH}^-]$

Key Diagrams

- Titration curve
- Titration curve

Points of Interest

Acid rain is a common name for acidic precipitation. It is rain or snow that has a pH of less than 7. It is caused by the presence of sulfur dioxide and nitrogen dioxide in the atmosphere.

Hydrogen ions are the smallest and most abundant ions in the environment. They are found in all acids and bases. They are also found in water and in the atmosphere.

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Key Equations

| | |
|--------------------------------------|--------------------------------------|
| $\text{pH} = -\log_{10}[\text{H}^+]$ | $\text{pH} = \log_{10}[\text{OH}^-]$ |
| $\text{pH} = 7$ | $\text{pH} = 7$ |
| $\text{pH} < 7$ | $\text{pH} > 7$ |
| $\text{pH} = 0$ | $\text{pH} = 14$ |
| $\text{pH} = 1$ | $\text{pH} = 13$ |
| $\text{pH} = 2$ | $\text{pH} = 12$ |
| $\text{pH} = 3$ | $\text{pH} = 11$ |
| $\text{pH} = 4$ | $\text{pH} = 10$ |
| $\text{pH} = 5$ | $\text{pH} = 9$ |
| $\text{pH} = 6$ | $\text{pH} = 8$ |
| $\text{pH} = 7$ | $\text{pH} = 7$ |
| $\text{pH} = 8$ | $\text{pH} = 6$ |
| $\text{pH} = 9$ | $\text{pH} = 5$ |
| $\text{pH} = 10$ | $\text{pH} = 4$ |
| $\text{pH} = 11$ | $\text{pH} = 3$ |
| $\text{pH} = 12$ | $\text{pH} = 2$ |
| $\text{pH} = 13$ | $\text{pH} = 1$ |
| $\text{pH} = 14$ | $\text{pH} = 0$ |

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