

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of acids and bases and vice versa.
2. Calculate pH and concentrations based on K_a and K_b and vice versa.
3. Calculate the pH of weak acids.

Vocabulary

- | | | |
|-------------------|-----------------------------|--------------------|
| 1. autoionization | 2. conjugate acid-base pair | 3. weak acid/base |
| 4. autoionization | 5. acid rain | 6. buffer solution |
| 7. autoionization | 8. acid rain | 9. pH |

Key Equation

- | | |
|------------------------------------|------------------------------------|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}^+]$ |

Part A Completion

Write the equation and/or the numerical value(s) of the quantity indicated that are indicated in the problem. Be sure to include units where appropriate.

1. What substance is _____, a weak hydrogen ion donor?
2. What substance is _____, the conjugate base of _____?
3. What substance is _____, a weak hydrogen ion acceptor?
4. What substance is _____, a weak hydrogen ion acceptor?
5. What substance is _____, a weak hydrogen ion acceptor?
6. What substance is _____, a weak hydrogen ion acceptor?
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19. What substance is _____, a weak hydrogen ion acceptor?
20. What substance is _____, a weak hydrogen ion acceptor?

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