

1-4 Study Guide and Intervention**The Distributive Property**

Evaluate Expressions The Distributive Property can be used to help evaluate expressions.

Distributive Property

For any numbers a , b , and c , $a(b + c) = ab + ac$ and $(b + c)a = ba + ca$ and $a(b - c) = ab - ac$ and $(b - c)a = ba - ca$.

Example 1 Use the Distributive Property to rewrite $6(8 + 10)$. Then evaluate.

$$\begin{aligned} 6(8 + 10) &= 6 \cdot 8 + 6 \cdot 10 && \text{Distributive Property} \\ &= 48 + 60 && \text{Multiply.} \\ &= 108 && \text{Add.} \end{aligned}$$

Example 2 Use the Distributive Property to rewrite $-2(3x^2 + 5x + 1)$. Then simplify.

$$\begin{aligned} -2(3x^2 + 5x + 1) &= -2(3x^2) + (-2)(5x) + (-2)(1) && \text{Distributive Property} \\ &= -6x^2 + (-10x) + (-2) && \text{Multiply.} \\ &= -6x^2 - 10x - 2 && \text{Simplify.} \end{aligned}$$

Exercises

Use the Distributive Property to rewrite each expression. Then evaluate.

1. $20(31)$ **620**

2. $12 \cdot 4 \frac{1}{2}$ **54**

3. $5(311)$ **1555**

4. $5(4x - 9)$ **$20x - 45$**

5. $3(8 - 2x)$ **$24 - 6x$**

6. $12\left(6 - \frac{1}{2}x\right)$ **$72 - 6x$**

7. $12\left(2 + \frac{1}{2}x\right)$ **$24 + 6x$**

8. $\frac{1}{4}(12 - 4x)$ **$3 - x$**

9. $3(2x - y)$ **$6x - 3y$**

10. $2(3x + 2y - z)$
 $6x + 4y - 2z$

11. $(x - 2)y$
 $xy - 2y$

12. $2(3x - 2b + c)$
 $6a - 4b + 2c$

13. $\frac{1}{4}(16x - 12y + 4z)$
 $4x - 3y + z$

14. $(2 - 3x + x^2)3$
 $6 - 9x + 3x^2$

15. $-2(2x^2 + 3x + 1)$
 $-4x^2 - 6x - 2$

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web solution total mass of the answer 100 g mass of benzene 30 g therefore mass of carbon tet 100 30 g benzene s molar mass C_6H_6 6 12 6 1 g mol 1 78 g mol 1 hence C_6H_6 s number of moles 30 78 mol 0 3846 mol molar mass of carbon tet CCl_4 1 12 4 355 154 g mol 1

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web jan 30 2020 ans the molality of solution is 1 mol kg 1 or 1 m example 02 11 11 g of urea nh_2CONH_2 was dissolved in 100 g of water calculate the molarity and molality of the solution given n 14 h 1 c 12 o 16 given mass of solute urea 11 11 g mass of solvent water 100 g 0 1 kg to find molarity of solution solution

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web solution moles of CH_3OH 5 5 32 0 17 mole moles of H_2O 40 18 2 2 moles therefore according to the equation mole fraction of CH_3OH 0 17 2 2 0 17 mole fraction of CH_3OH 0 073 to solve more examples on mole fraction formula and practice more question please visit byju s com

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal KBr that contains 2 92 kg of water 1 molality moles solute divided by kilograms solute 0 449 mol kg x 2 92 kg x 1 31108 mol of KBr 2 moles times molar mass equals grams 1 31108 mol 119 0023 g mol 156 g KBr

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web calculate the mass percent mole fraction molality and molarity of hexane c 6 h 14 d 0 660 g cm 3 when it is mixed with

heptane c 7 h 12 d 0 684 g cm³ in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web jan 8 2022 mol xa x 100 the sum of all of the mole percents of a mixture equals 100 mole fraction properties and advantages the mole fraction offers advantages over some of the other units of concentration unlike molarity mole fraction is

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web converting moles and mass the molecular weight of sodium chloride nacl is 58 44 g mol how many moles of salt are in 13 8 g of sodium chloride express the answer using 3 significant figures stuck use a hint learn for free about math art computer programming economics physics chemistry biology medicine finance history and more

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web a derivatization that achieves this goes as molarity moles of solute volume of solution in liter moles molarity x volume in liter mass moles x molar mass numerical answers 1

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web what is the mole fraction of licl in the solution 11 8 practice problem acetone c 3 h 6 o has a density of 0 791 g ml and xylene c 8 h 10 has a density of 0 864 g ml a solution was prepared by mixing acetone and xylene calculate the mole fraction of acetone in a solution if 19 0 ml of acetone was dissolved in 98 4 ml of xylene 7 1

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web since the mole fraction of the O_2 is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.21. 0.050 = 1.05 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature