

1-4 Study Guide and Intervention***The Distributive Property***

Evaluate Expressions The Distributive Property can be used to help evaluate expressions.

Distributive Property

For any numbers a , b , and c , $a(b + c) = ab + ac$ and $(b + c)a = ba + ca$ and $a(b - c) = ab - ac$ and $(b - c)a = ba - ca$.

Example 1

Use the Distributive Property to rewrite $6(8 + 10)$. Then evaluate.

$$\begin{aligned} 6(8 + 10) &= 6 \cdot 8 + 6 \cdot 10 \\ &= 48 + 60 \\ &= 108 \end{aligned}$$

Distributive Property

Multiply.

Add.

Example 2

Use the Distributive Property to rewrite $-2(3x^2 + 5x + 1)$. Then simplify.

$$\begin{aligned} -2(3x^2 + 5x + 1) &= -2(3x^2) + (-2)(5x) + (-2)(1) && \text{Distributive Property} \\ &= -6x^2 + (-10x) + (-2) && \text{Multiply.} \\ &= -6x^2 - 10x - 2 && \text{Simplify.} \end{aligned}$$

Exercises

Use the Distributive Property to rewrite each expression. Then evaluate.

1. $20(31)$ **620**

2. $12 \cdot 4\frac{1}{2}$ **54**

3. $5(311)$ **1555**

4. $5(4x - 9)$ **$20x - 45$**

5. $3(8 - 2x)$ **$24 - 6x$**

6. $12\left(6 - \frac{1}{2}x\right)$ **$72 - 6x$**

7. $12\left(2 + \frac{1}{2}x\right)$ **$24 + 6x$**

8. $\frac{1}{4}(12 - 4f)$ **$3 - f$**

9. $3(2x - y)$ **$6x - 3y$**

10. $2(3x + 2y - z)$
 $6x + 4y - 2z$

11. $(x - 2)y$
 $xy - 2y$

12. $2(3a - 2b + c)$
 $6a - 4b + 2c$

13. $\frac{1}{4}(16x - 12y + 4z)$
 $4x - 3y + z$

14. $(2 - 3x + x^2)3$
 $6 - 9x + 3x^2$

15. $-2(2x^2 + 3x + 1)$
 $-4x^2 - 6x - 2$

Distributive Property Study Guide

CK-12 Foundation



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Manual, Mathematics for Elementary Teachers, a Contemporary Approach, Fourth Edition, Gary L. Musser, William S. [sic] Burger Don Miller, 1997 College Algebra, 4e Study Guide Miller, 1998-03-21 **Algebra 1 Study Guide** Miriam A. Leiva, 2001

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web nov 6 2023 x_a mol a mol a mol b and x_b mol b mol a mol b note in the given mixture the sum of all the mole fractions present is equal to one $x_a + x_b = 1$ when the mole fractions are multiplied by 100 they give the mole percentage mole fraction is a unitless and dimensionless expression image will be uploaded soon

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web solution total mass of the answer 100 g mass of benzene 30 g therefore mass of carbon tet 100 30 g benzene s molar mass C_6H_6 6 12 6 1 g mol 1 78 g mol 1 hence C_6H_6 s number of moles $\frac{30}{78}$ mol 0 3846 mol molar mass of carbon tet CCl_4 1 12 4 355 154 g mol 1

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web jan 30 2020 ans the molality of solution is 1 mol kg 1 or 1 m example 02 11 11 g of urea NH_2CONH_2 was dissolved in 100 g of water calculate the molarity and molality of the solution given n 14 h 1 c 12 o 16 given mass of solute urea 11 11 g mass of solvent water 100 g 0 1 kg to find molarity of solution solution

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web solution moles of CH_3OH 5 5 32 0 17 mole moles of H_2O 40 18 2 2 moles therefore according to the equation mole fraction of CH_3OH $\frac{0 17}{0 17 + 2 2}$ 0 073 to solve more examples on mole fraction formula and practice more question please visit byju s com

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web you may work out the mole fractions on your own problem 5 what is the mass of a sample of a 0 449 molal kbr that contains 2 92 kg of water 1 molality moles solute divided by kilograms solute $\frac{0 449 \text{ mol kg} \times 2 92 \text{ kg} \times 1}{31108 \text{ mol of kbr} \times 2}$ moles times molar mass equals grams $\frac{1 31108 \text{ mol} \times 119 0023 \text{ g mol}}{156 \text{ g kbr}}$

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web calculate the mass percent mole fraction molality and molarity of hexane C_6H_{14} d 0 660 g cm^{-3} when it is mixed with

heptane c 7 h 12 d 0 684 g cm³ in a 35 0 55 0 ml ratio assume that the volumes are additive

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0 2885 its mole percent would be 28 85

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web since the mole fraction of the O_2 is 0.050 it contributes 0.050 or if you will 5.0% of the total pressure thus total pressure is 0.2105042 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature