

Chemical Reactions:

1. The following reaction $\text{PbO}_2 + \text{H}_2\text{O}_2 \rightarrow \text{PbO}_2\text{O}_3 + \text{H}_2\text{O}$ is not true. Which is not true?

A) reaction with 2.0 mol/l H_2O_2
 B) ground reaction with 3.0 mol/l H_2O_2
 C) reaction with 1.0 mol/l H_2O_2
 D) ground reaction with 2.0 mol/l H_2O_2
 E) reaction with 4.0 mol/l H_2O_2

2. A 1.0 mol/l example of the NO_x reaction is oxygenated according to the equation: $2\text{NO}_2\text{O}_2(\text{g}) \rightarrow 2\text{NO}_2\text{O}(\text{g}) + 2\text{O}_2(\text{g})$. Below, which reaction of O_2 with NO_2 will result in the same amount of NO_2 gas?

A) 0.1 mol
 B) 1.0 mol
 C) 1.2 mol
 D) 0.7 mol
 E) 2.0 mol

3. The reaction of NO_2 is step characterized by the equation: $4\text{NO}_2 + \text{O}_2 \rightleftharpoons 2\text{NO}_2\text{O}_2$. If reaction 2.0 mol/l NO_2 and 0.5 mol/l O_2 is carried out, then which reaction will affect the reaction rate most strongly?

A) 0.5 mol/l
 B) 0.7 mol/l
 C) 1.0 mol/l
 D) 1.5 mol/l
 E) 2.0 mol/l

4. Calculated the following reaction, v from 2000 mol/l NO_2 and 2000 mol/l O_2 : $2\text{NO}_2\text{O}_2(\text{g}) \rightarrow 2\text{NO}_2(\text{g}) + 2\text{O}_2(\text{g})$

$\text{v} = \text{P}_2\text{O}_5 \text{ g/mol} \times 20 \text{ mol/l} \times \text{v}$ reaction. In the reaction, the difference between the production of 4.0 mol/l of NO_2 , which is the difference between 2000 mol/l NO_2 and 2000 mol/l O_2 .

A) 0.6
 B) 0.8
 C) 0.9
 D) 1.0
 E) 1.2

5. Under the following conditions, v $\text{NO}_2\text{O}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons 2\text{NO}_2\text{O}(\text{g}) + \text{O}_2(\text{g})$ is 2000 mol/l NO_2 and 2000 mol/l O_2 . How many molecules of NO_2 are present in the reaction?

A) 1.00 $\times 10^{22}$
 B) 2.00 $\times 10^{22}$
 C) 4.00 $\times 10^{22}$
 D) 8.00 $\times 10^{22}$
 E) 16.00 $\times 10^{22}$

6. The following reaction $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{H}_2\text{O}(\text{g})$ is carried out. Below, which reaction of H_2 and O_2 will the ground reaction between 1.00 mol/l H_2 and 1.00 mol/l O_2 result?

A) 1.00 mol/l
 B) 2.00 mol/l
 C) 0.500 mol/l
 D) 0.100 mol/l
 E) 0.200 mol/l

7. Under the following conditions of reaction, which reaction of O_2 with CH_3O_2 is oxygenated to react, incompletely with 1.00 g of CH_3O_2 ?

A) 0.000 $\times 10^{22}$ g
 B) 0.001 $\times 10^{22}$ g
 C) 0.005 $\times 10^{22}$ g
 D) 0.010 $\times 10^{22}$ g
 E) 0.100 $\times 10^{22}$ g

8. Under reaction conditions and quantities of reaction, which reaction of O_2 with CH_3O_2 is oxygenated to react, incompletely with 1.00 g of CH_3O_2 ?

$\text{C}_2\text{H}_5\text{O}_2(\text{g}) + \text{O}_2 \rightleftharpoons 2\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$

A) 1.00 $\times 10^{22}$ molecules
 B) 1.00 $\times 10^{23}$ molecules
 C) 4.00 $\times 10^{22}$ molecules
 D) 8.00 $\times 10^{22}$ molecules
 E) 16.00 $\times 10^{22}$ molecules

9. Calculate the mass of carbon dioxide produced from 2000 mol/l CO_2 in the following reaction.

$2\text{C}_2\text{H}_5\text{O}_2(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow 2\text{CO}_2(\text{g}) + 2\text{H}_2\text{O}(\text{g})$

A) 7.00 $\times 10^{22}$ g CO_2
 B) 1.400 $\times 10^{22}$ g CO_2
 C) 2.80 $\times 10^{22}$ g CO_2
 D) 4.200 $\times 10^{22}$ g CO_2
 E) 1.800 $\times 10^{22}$ g CO_2

10. Calculate the mass of oxygen required to react with 1.00 g of sulfur in the following reaction.

$2\text{S}(\text{s}) + \text{O}_2(\text{g}) \rightarrow 2\text{SO}_2(\text{g})$

A) 1.000 $\times 10^{22}$ molecules O_2
 B) 0.500 $\times 10^{22}$ molecules O_2
 C) 1.75 $\times 10^{22}$ molecules O_2
 D) 3.50 $\times 10^{22}$ molecules O_2
 E) 0.250 $\times 10^{22}$ molecules O_2

Multiple Choice Stoichiometry Test

Sasha Alcott

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