

## KEY

AP CHEMISTRY MULTIPLE CHOICE QUESTIONS  
Unit 8 – Introduction to Equilibrium

The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the equilibrium constant for the following reaction at 25°C?



(A) 95

(B) 95

(C)  $95^2$ (D)  $\frac{1}{95}$ (E)  $\frac{1}{(95)^2}$ 

2. For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products? *increase "P"*



Shift towards fewer gas

4.



When 4.00 mol each of  $X(g)$  and  $Y(g)$  are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of  $Z(g)$  is produced. What is the value of the equilibrium constant,  $K_c$ ?

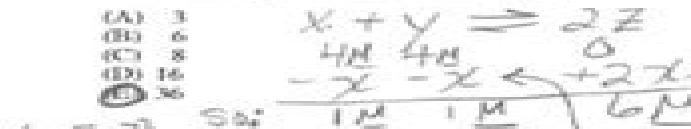
(A) 3

(B) 6

(C) 8

(D) 16

(E) 36



6.

The synthesis of  $CH_3OH(g)$  from  $CO(g)$  and  $H_2(g)$  is represented by the equation above. The value of  $K_c$  for the reaction at 483 K is 14.8.

*increase heat shifts left, decrease K*

Which of the following explains the effect on the equilibrium constant,  $K_c$ , when the temperature of the reaction system is increased to 650 K?

(A)  $K_c$  will increase because the activation energy of the forward reaction increases more than that of the reverse reaction.

(B)  $K_c$  will increase because there are more reactant molecules than product molecules.

(C)  $K_c$  will decrease because the reaction is exothermic.

(D)  $K_c$  is constant and will not change.



$C_2H_6(g)$  is reduced by  $H_2(g)$  in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- Changes in the partial pressure of  $H_2(g)$
- Changes in the particle size of the platinum catalyst
- Changes in the temperature of the reaction system

(A) III only  
(B) I and II only  
(C) I and III only  
(D) II and III only  
(E) I, II, and III

*← Needs to shift left*  
shifts  $2 BaO_2(s) \rightleftharpoons 2 BaO(s) + O_2(g)$

$$\Delta H^\circ = 162 \text{ kJ/mol}_\text{rxn} \text{ endothermic}$$

A sealed rigid vessel contains  $BaO_2(s)$  in equilibrium with  $BaO(s)$  and  $O_2(g)$  as represented by the equation above. Which of the following changes will increase the amount of  $BaO_2(s)$  in the vessel?

(A) Removing a small amount of  $O_2(g)$  *R*  
(B) Removing a small amount of  $BaO(s)$  *X*  
(C) Adding He gas to the vessel *X*  
(D) Lowering the temperature *L*

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