

## 14.3

## HYDROGEN IONS AND ACIDITY

## Section Review

## Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentrations from pH and hydroxide ion concentrations from pOH.
3. Calculate the pH of a weak acid.

## Vocabulary

- |                   |                               |                     |
|-------------------|-------------------------------|---------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution    |
| 4. autoionization | 5. pOH                        | 6. neutral solution |

## Key Equation

- |                                     |                                     |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

## Part A Completion

Write the hydrogen ion concentration or hydroxide ion concentration of the solution described. Round the concentration to the correct significant figures. Express acids and bases in scientific notation.

1. What substance is \_\_\_\_\_, contains hydrogen ions?
2. \_\_\_\_\_ is the concentration of those acids present in water and equal to \_\_\_\_\_.
3. \_\_\_\_\_ is the concentration of those acids present in water and equal to \_\_\_\_\_.
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# 19 2 Hydrogen Ions And Acidity Worksheet Answers

**Advisory Council on College Chemistry**

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