

ACIDS, BASES, & THE pH SCALE

Name: *Answer Key*

Date:

Pd:

PART 1: ACIDS AND BASES

Comparing Acids and Bases	
Acids	Bases (Alkalines)
• Sour taste, leaves burning sensation	• Sharp bitter taste, feels slippery to touch
• Increase hydrogen ions (H^+) in solution	• Decrease hydrogen ions (H^+) in solution (Increase hydroxide ions (OH^-) in solution)
• Example: HCl $H^+ + Cl^-$	• Example: $NaOH$ $Na^+ + OH^-$
• Turns litmus (pH) paper to red	• Turns litmus (pH) paper to blue

PART 2: THE pH SCALE

- The pH Scale is a measurement of how acidic or how basic a solution is.
- The scale goes from 0 to 14.
- Compounds are acidic if they have a pH lower than 7.
- Compounds with a pH higher than 7 are basic or alkaline.
- Compounds with a pH of 7 are neutral.



Increased concentration of H^+ ions

Decreased concentration of OH^- ions

pH below 7 = acids

pH at 7 = neutral
(pure water)

Decreased concentration of H^+ ions

Increased concentration of OH^- ions

pH above 7 = bases

PART 3: BUFFERS

- A buffer is a weak acid or base that prevents sudden changes in pH.
 - Maintains homeostasis
- A buffer works by binding to H^+ ions releasing H^+ ions when the solution pH changes.
- Buffers bind to H^+ ions when the concentration of H^+ ions in solution increases.
- $NaOH$ $Na^+ + OH^-$ (The OH^- will bind to H^+ ions to form H_2O – neutralization)
- Buffers release H^+ ions when the concentration of H^+ ions in solution decreases.
 HCl $H^+ + Cl^-$

Acids And Bases Study Guide Key

T. W. Graham Solomons

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Study Guide with Selected Solutions Karen Timberlake,2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems

Study Guide with Answers to Selected Problems H. Stephen Stoker,Danny V. White,2003-04 This supplement includes for each chapter a brief overview activities and practice problems to reinforce skills and a practice test The answers section includes answers for all odd numbered end of chapter exercises

Fundamentals of Organic Chemistry, Textbook, Study Guide and Solutions Manual T. W. Graham Solomons,1996-08-15 A realistic approach to the study of mechanisms The book addresses real functional group chemistry with an emphasis on the biological environmental and medical applications of organic chemistry

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pH buffering system Mário Sérgio Piantavini,2025-09-19 Acid Base Balance A Comprehensive Study Guide explores the fundamental principles of pH regulation and buffer systems in human physiology The book presents a historical overview of acid base research from the early 20th century to modern developments followed by an in depth study of key buffering mechanisms the role of lungs and kidneys in homeostasis and the diagnostic use of blood gas analysis It examines metabolic and respiratory disorders their compensatory responses and the complexities of mixed acid base disturbances Special clinical contexts such as cirrhosis hemodialysis and neonatal physiology are addressed highlighting the challenges of maintaining acid base stability in vulnerable populations The book also introduces modern analytical technologies mathematical models and innovative monitoring systems bridging theory and clinical application This guide serves as a valuable resource for students healthcare professionals and researchers seeking a clear yet thorough understanding of acid base physiology and its clinical implications **Study Guide for Introductory Chemistry, a Foundation, Introductory Chemistry, Basic Chemistry** Iris Stovall,Steven S. Zumdahl,1996 **General Chemistry Study Guide Sixth Edition** Darrell D. Ebbing,1999

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103 d 2 1 102 e 2 1 103 2 a sample of helium gas occupies 2 65 l at 1 20 atm what pressure would this sample of gas exert in a 1 50 l container at the same temperature

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