

ACIDS, BASES, & THE pH SCALE

Name: Answer Key

Date: _____

Pd: _____

PART 1: ACIDS AND BASES

Comparing Acids and Bases	
Acids	Bases (Alkalines)
<ul style="list-style-type: none">• <u>Sour</u> taste, leaves <u>burning</u> sensation• <u>Increase</u> hydrogen ions (H^+) in solution• Example: <u>HCl</u> <u>$H^+ + Cl^-$</u>• Turns <u>litmus</u> (pH) paper to <u>red</u>	<ul style="list-style-type: none">• Sharp <u>bitter</u> taste, feels <u>slippery</u> to touch• <u>Decrease</u> hydrogen ions (H^+) in solution (<u>Increase</u> hydroxide ions (OH^-) in solution)• Example: <u>$NaOH$</u> <u>$Na^+ + OH^-$</u>• Turns litmus (pH) paper to <u>blue</u>

PART 2: THE pH SCALE

- The pH Scale is a measurement of how acidic or how basic a solution is.
- The scale goes from 0 to 14.
- Compounds are acidic if they have a pH lower than 7.
- Compounds with a pH higher than 7 are basic or alkaline.
- Compounds with a pH of 7, are neutral.

0 1 2 3 4 5 6 7 8 9 10 11 12 13 14

Increased concentration of H^+ ions
Decreased concentration of OH^- ions

Decreased concentration of H^+ ions
Increased concentration of OH^- ions

pH below 7 = acids

pH at 7 = neutral
(pure water)

pH above 7 = bases

PART 3: BUFFERS

- A buffer is a weak acid or base that prevents sudden changes in pH.
 - Maintains homeostasis
- A buffer works by binding to H^+ ions releasing H^+ ions when the solution pH changes.
- Buffers bind to H^+ ions when the concentration of H^+ ions in solution increases.
- $NaOH$ $Na^+ + OH^-$ (The OH^- will bind to H^+ ions to form H_2O - neutralization)
- Buffers release H^+ ions when the concentration of H^+ ions in solution decreases.
 HCl $H^+ + Cl^-$

Acids And Bases Study Guide Key

T. W. Graham Solomons



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