

ACIDS, BASES AND SALTS

Definition of an acid An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	What causes the acidic properties of acids? The hydrogen ions (H^+) cause the acidic properties and these are formed in the presence of water.
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Another term to refer to an acid An acid is called a <u>proton donor</u>	Why an acid is also called a proton donor? It's because an acid provides protons or hydrogen ions (H^+) to other substances during the reaction.
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Substance to which an acid provides protons BASE	A base is then called PROTON ACCEPTOR	Why the base is called a proton acceptor? Because the base accepts hydrogen ions from acids	Equation for the reaction between acid and base $H^+_{(aq)} + OH^-_{(aq)} \longrightarrow H_2O_{(l)}$
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Common laboratory acids		These three common laboratory acids are also called	Why these acids are also called mineral acids?
Hydrochloric acid	HCl	MINERAL ACIDS	They are derived from mineral salts i.e chlorides for HCl, sulphates for H_2SO_4 and nitrates for HNO_3
Sulphuric acid	H_2SO_4		
Nitric acid	HNO_3		

Other mineral acids known		Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid	H_2SO_3	Derived from sulphites	Ethanoic acid (CH_3COOH)	CITRIC ACID from lemons
Carbonic acid	H_2CO_3	Derived from carbonates		TARTARIC ACID from grapes
Phosphoric acid	H_3PO_4	Derived from phosphates	Methanoic acid ($HCOOH$)	ACETIC ACID from vinegar
Nitrous acid	HNO_2	Derived from nitrites		LACTIC ACID from sour milk
				Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces HYDROGEN IONS	Term given to the number of hydrogen ions produced by one molecule of an acid BASICITY OF AN ACID	Definition of basicity of an acid BASICITY of an acid is the <u>number of hydrogen ions</u> produced by <u>one molecule</u> of an acid <u>in aqueous solution</u> .
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Acid And Bases Study Guide

Michael Seilmaier



Acid And Bases Study Guide:

Acid-Base Equilibria - Quick Chemistry Review Outline and Handout E Staff, Acid Base Equilibria Quick Review Outline and Handout for All Students Learn and review on the go Use Quick Review Chemistry Notes to help you learn or brush up on the subject quickly You can use the review notes as a reference to understand the subject better and improve your grades Easy to remember facts to help you perform better Perfect study notes for all high school and college students 10 Pages

Organic Chemistry, 12e Binder Ready Version Study Guide & Student Solutions Manual T. W. Graham Solomons, Craig B. Fryhle, Scott A. Snyder, 2016-04-11 This is the Student Study Guide Solutions Manual to accompany Organic Chemistry 12th Edition The 12th edition of Organic Chemistry continues Solomons Fryhle Snyder s tradition of excellence in teaching and preparing students for success in the organic classroom and beyond A central theme of the authors approach to organic chemistry is to emphasize the relationship between structure and reactivity To accomplish this the content is organized in a way that combines the most useful features of a functional group approach with one largely based on reaction mechanisms The authors philosophy is to emphasize mechanisms and their common aspects as often as possible and at the same time use the unifying features of functional groups as the basis for most chapters The structural aspects of the authors approach show students what organic chemistry is Mechanistic aspects of their approach show students how it works And wherever an opportunity arises the authors show students what it does in living systems and the physical world around us

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

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Study Guide and Solutions Manual, Fundamentals of General, Organic, and Biological Chemistry, Third Edition John McMurry, Susan McMurry, 1999 Provides worked out solutions to text problems along with chapter by chapter outlines and a variety of self tests at the end of each chapter

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Table of Contents Acid And Bases Study Guide

1. Understanding the eBook Acid And Bases Study Guide
 - The Rise of Digital Reading Acid And Bases Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid And Bases Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid And Bases Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid And Bases Study Guide
 - Personalized Recommendations
 - Acid And Bases Study Guide User Reviews and Ratings
 - Acid And Bases Study Guide and Bestseller Lists
5. Accessing Acid And Bases Study Guide Free and Paid eBooks

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- Acid And Bases Study Guide eBook Subscription Services
- Acid And Bases Study Guide Budget-Friendly Options
- 6. Navigating Acid And Bases Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Acid And Bases Study Guide Compatibility with Devices
 - Acid And Bases Study Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid And Bases Study Guide
 - Highlighting and Note-Taking Acid And Bases Study Guide
 - Interactive Elements Acid And Bases Study Guide
- 8. Staying Engaged with Acid And Bases Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid And Bases Study Guide
- 9. Balancing eBooks and Physical Books Acid And Bases Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid And Bases Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acid And Bases Study Guide
 - Setting Reading Goals Acid And Bases Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acid And Bases Study Guide
 - Fact-Checking eBook Content of Acid And Bases Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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