

CHAPTER 14

ACIDS AND BASES

Questions

19. Acids are proton (H^+) donors, and bases are proton acceptors.

HCO_3^- as an acid: $\text{HCO}_3^-(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{CO}_3^{2-}(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$

HCO_3^- as a base: $\text{HCO}_3^-(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_2\text{CO}_3(\text{aq}) + \text{OH}^-(\text{aq})$

H_2PO_4^- as an acid: $\text{H}_2\text{PO}_4^- + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{HPO}_4^{2-}(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$

H_2PO_4^- as a base: $\text{H}_2\text{PO}_4^- + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{PO}_4(\text{aq}) + \text{OH}^-(\text{aq})$

20. Acidic solutions (at 25°C) have an $[\text{H}^+] > 1.0 \times 10^{-7} \text{ M}$, which gives a $\text{pH} < 7.0$. Because $[\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ and $\text{pH} + \text{pOH} = 14.00$ for an aqueous solution at 25°C, an acidic solution must also have $[\text{OH}^-] < 1.0 \times 10^{-7} \text{ M}$ and $\text{pOH} > 7.00$. From these relationships, the solutions in parts a, b, and d are acidic. The solution in part c will have a $\text{pH} > 7.0$ ($\text{pH} = 14.00 - 4.51 = 9.49$) and is therefore not acidic (solution is basic).

21. Basic solutions (at 25°C) have an $[\text{OH}^-] > 1.0 \times 10^{-7} \text{ M}$, which gives a $\text{pOH} < 7.0$. Because $[\text{H}^+][\text{OH}^-] = 1.0 \times 10^{-14}$ and $\text{pH} + \text{pOH} = 14.00$ for any aqueous solution at 25°C, a basic solution must also have $[\text{H}^+] < 1.0 \times 10^{-7} \text{ M}$ and $\text{pH} > 7.00$. From these relationships, the solutions in parts b, c, and d are basic solutions. The solution in part a will have a $\text{pH} < 7.0$ ($\text{pH} = 14.00 - 11.21 = 2.79$) and is therefore not basic (solution is acidic).

22. When a strong acid (HX) is added to water, the reaction $\text{HX} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{X}^-$ essentially goes to completion. All strong acids in water are completely converted into H_3O^+ and X^- . Thus no acid stronger than H_3O^+ will remain undissociated in water. Similarly, when a strong base (B) is added to water, the reaction $\text{B} + \text{H}_2\text{O} \rightarrow \text{BH}^+ + \text{OH}^-$ essentially goes to completion. All bases stronger than OH^- are completely converted into OH^- and BH^+ . Even though there are acids and bases stronger than H_3O^+ and OH^- , in water these acids and bases are completely converted into H_3O^+ and OH^- .

23. 10.78 (4 S.F.); 6.78 (3 S.F.); 0.78 (2 S.F.); a pH value is a logarithm. The numbers to the left of the decimal point identify the power of 10 to which $[\text{H}^+]$ is expressed in scientific notation, for example, 10^{-11} , 10^{-7} , 10^{-1} . The number of decimal places in a pH value identifies the number of significant figures in $[\text{H}^+]$. In all three pH values, the $[\text{H}^+]$ should be expressed only to two significant figures because these pH values have only two decimal places.

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