

ACIDS, BASES AND SALTS

Definition of an acid An acid is a compound which when dissolved in water produces hydrogen ions as the only positively charged ions	What causes the acidic properties of acids? The hydrogen ions (H^+) cause the acidic properties and these are formed in the presence of water.
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Another term to refer to an acid An acid is called a proton donor	Why an acid is also called a proton donor? It's because an acid provides protons or hydrogen ions (H^+) to other substances during the reaction.
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Substance to which an acid provides protons BASE	A base is then called PROTON ACCEPTOR	Why the base is called a proton acceptor? Because the base accepts hydrogen ions from acids	Equation for the reaction between acid and base $H^{+}_{(aq)} + OH^{-}_{(aq)} \longrightarrow H_2O_{(l)}$
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Common laboratory acids		These three common laboratory acids are also called MINERAL ACIDS	Why these acids are also called mineral acids? They are derived from mineral salts ie chlorides for HCl , sulphates for H_2SO_4 and nitrates for HNO_3
Hydrochloric acid	HCl		
Sulphuric acid	H_2SO_4		
Nitric acid	HNO_3		

Other mineral acids known	Mineral salts from which the acid is derived	Organic acids known	Naturally occurring acids
Sulphurous acid	H_2SO_3	Derived from sulphites	CITRIC ACID from lemons
Carbonic acid	H_2CO_3	Derived from carbonates	TARTARIC ACID from grapes
Phosphoric acid	H_3PO_4	Derived from phosphates	ACETIC ACID from vinegar
Nitrous acid	HNO_2	Derived from nitrates	LACTIC ACID from sour milk
		Methanoic acid ($HCOOH$)	Hydrochloric acid from digestive juices

Whenever an acid is dissolved in water, it produces HYDROGEN IONS	Term given to the number of hydrogen ions produced by one molecule of an acid BASICITY OF AN ACID	Definition of basicity of an acid BASICITY of an acid is the <u>number of hydrogen ions produced by one molecule of an acid in aqueous solution</u> .
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Acid Bases Study Guide

James E. Brady, John R. Holum

Acid Bases Study Guide:

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