

REPORT SHEET | LAB
Acid-Base Titration | 10

A. Concentration of Acetic Acid in Vinegar

1. Brand Heinz Volume 5.00 mL (% on label) 5.0 %
 2. Molarity (M) of NaOH _____ M

	Trial 1	Trial 2	Trial 3
3. Initial NaOH level in buret	0.1 mL	0.2 mL	0.1 mL
4. Final NaOH level in buret	39.6 mL	41.2 mL	40.1 mL
5. Volume (mL) of NaOH used			
6. Average volume (mL)			
7. Average volume in liters (L)			

Show all calculations for the following questions.

8. Moles of NaOH used in titration _____ mole NaOH
9. Moles of $\text{HC}_2\text{H}_3\text{O}_2$ neutralized by NaOH _____ mole $\text{HC}_2\text{H}_3\text{O}_2$
10. Molarity of $\text{HC}_2\text{H}_3\text{O}_2$ _____ M $\text{HC}_2\text{H}_3\text{O}_2$
11. Grams of $\text{HC}_2\text{H}_3\text{O}_2$ _____ g $\text{HC}_2\text{H}_3\text{O}_2$
12. Percent (m/v) $\text{HC}_2\text{H}_3\text{O}_2$ in vinegar _____ % $\text{HC}_2\text{H}_3\text{O}_2$

Acid Base Titration Lab Chemfax Answers

Robert De Levie

Acid Base Titration Lab Chemfax Answers:

Acid-base Titrations in Nonaqueous Solvents James Sherwood Fritz, 1952 **An On-line Acid-base Titration Applet in the Generic Tutorial System for the Sciences Project** Thomas Lee Gummo, 2002 The purpose of this Master's Project was to develop an Acid Base Titration Simulator. It was also to be a part of the California State University San Bernardino's GTSS Generic Tutorial System for the Sciences project. The main benefit is that students will be able to conduct titration experiments over the Internet without being in the laboratory and without costly equipment or dangerous chemicals.

Instructors at the high school and college level can demonstrate the key principles of titration **EDTA Titrations** H. A. Flaschka, 2013-10-22 EDTA Titrations An Introduction to Theory and Practice Second Edition considers the theoretical background full procedural details and some practical applications of EDTA titrations. Ethylenediamine tetra acetic acid EDTA has risen from an obscure chemical compound to the most widely used organic reagent. This book is composed of 21 chapters. The opening chapters present the general theoretical foundations of EDTA titrations. The subsequent chapters describe the properties of EDTA such as the stability constants, titration curves, selectivity and masking effect. These topics are followed by discussions on titration types, standard solutions and reagents. The remaining chapters cover some of the practical applications of EDTA titrations. This book is directed toward students with advanced courses in analytical and organic chemistry *Working with Chem Separate* Wink, 1999-12-01 **Aqueous Acid-base Equilibria and Titrations**

Robert De Levie, 1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry. It is in addition a fresh and modern approach to a topic all too often taught in an outmoded way. This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems; these approaches are often simpler than the approximations which they replace. Acid base problems are readily addressed in terms of the proton condition, a convenient amalgam of the mass and charge constraints of the chemical system considered. The graphical approach of Bjerrum, Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria. Based on these concentrations, the proton condition can usually be simplified, often leading directly to the value of the pH. In the description of acid base titrations a general master equation is developed. It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data. Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated. Activity effects are described in detail including their effect on titration curves. The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not. Finally an entire chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved.

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