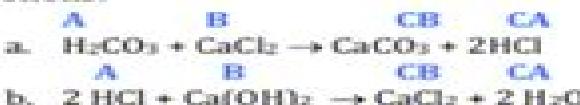


Chapter 19 Study Guide: Acids, Bases & Titration – Answers

1. Differentiate between Arrhenius, Brønsted-Lowry & Lewis acids and bases.

	ACIDS	BASES
Arrhenius	Form hydronium ions (H_3O^+) in aqueous sol'n	Form hydroxide ions (OH^-) in aqueous sol'n
Brønsted-Lowry	Proton (H^+) donors	Proton (H^+) acceptors
Lewis	Electron pair acceptors	Electron pair donors

2. Label the acid (A), base (B), conjugate acid (CA) and conjugate base (CB) in each of the following reactions:



3. Write the names or formulas for the following acids and bases:

a. H_2S – hydrosulfuric acid	e. Nitrous acid – HNO_2
b. $\text{HC}_2\text{H}_3\text{O}_2$ – acetic acid	f. Calcium hydroxide – $\text{Ca}(\text{OH})_2$
c. NaOH – sodium hydroxide	g. hydrochloric acid – HCl
d. HNO_3 – nitric acid	

Use the following formulas for the remaining questions: (these will not be given to you on the test!)

$$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.0 \times 10^{-14}$$

$$\text{pH} = -\log[\text{H}_3\text{O}^+]$$

$$\text{pOH} = -\log[\text{OH}^-]$$

$$\text{pH} + \text{pOH} = 14$$

$$\text{M}_\text{A}\text{-Va}^- \text{M}_\text{B} = \text{M}_\text{B}\text{-V}_\text{a}^- \text{M}_\text{A}$$

4. What is the pH of 1.34×10^{-4} M solution of hydrochloric acid?

$$\text{pH} = -\log[1.34 \times 10^{-4}\text{M}] = 3.873$$

5. Determine the pH and pOH of a 1.85×10^{-6} M NaOH solution.

$$\text{pOH} = -\log[1.85 \times 10^{-6}\text{M}] = 5.733$$

$$\text{pH} = 14 - 5.73 = 8.267$$

6. What is the $[\text{OH}^-]$ of an aqueous solution that has a pH of 5.0?

$$\text{pOH} = 9 \quad [\text{OH}^-] = 1 \times 10^{-9}\text{M}$$

7. What is the pH and the pOH of a solution with a volume of 2000. mL containing 35 grams of NaOH? Is this solution acidic or basic?

$$\frac{35\text{gNaOH}}{2.000\text{L}} \times \frac{1\text{molNaOH}}{40.00\text{gNaOH}} = 0.44\text{M NaOH}$$

$$\text{pOH} = -\log[0.44\text{M}] = 0.36 \quad \text{pH} = 14 - 0.36 = 13.64 \quad \text{basic}$$

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Pamela Mills,Amina El-Ashmawy

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Table of Contents Acid And Bases Study Guide Chap 15

1. Understanding the eBook Acid And Bases Study Guide Chap 15
 - The Rise of Digital Reading Acid And Bases Study Guide Chap 15
 - Advantages of eBooks Over Traditional Books
2. Identifying Acid And Bases Study Guide Chap 15
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acid And Bases Study Guide Chap 15
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acid And Bases Study Guide Chap 15
 - Personalized Recommendations
 - Acid And Bases Study Guide Chap 15 User Reviews and Ratings

- Acid And Bases Study Guide Chap 15 and Bestseller Lists
- 5. Accessing Acid And Bases Study Guide Chap 15 Free and Paid eBooks
 - Acid And Bases Study Guide Chap 15 Public Domain eBooks
 - Acid And Bases Study Guide Chap 15 eBook Subscription Services
 - Acid And Bases Study Guide Chap 15 Budget-Friendly Options
- 6. Navigating Acid And Bases Study Guide Chap 15 eBook Formats
 - ePUB, PDF, MOBI, and More
 - Acid And Bases Study Guide Chap 15 Compatibility with Devices
 - Acid And Bases Study Guide Chap 15 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acid And Bases Study Guide Chap 15
 - Highlighting and Note-Taking Acid And Bases Study Guide Chap 15
 - Interactive Elements Acid And Bases Study Guide Chap 15
- 8. Staying Engaged with Acid And Bases Study Guide Chap 15
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acid And Bases Study Guide Chap 15
- 9. Balancing eBooks and Physical Books Acid And Bases Study Guide Chap 15
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acid And Bases Study Guide Chap 15
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acid And Bases Study Guide Chap 15
 - Setting Reading Goals Acid And Bases Study Guide Chap 15
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acid And Bases Study Guide Chap 15
 - Fact-Checking eBook Content of Acid And Bases Study Guide Chap 15
 - Distinguishing Credible Sources

- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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