

Chemistry: Acids and Bases Study Guide (Ch. 15)

1. Give the definitions of an acid and a base according to the Arrhenius' Definition.
2. Give the definition of an acid and a base according to the bronsted-lowry definition.
3. Which of the definitions is used more often? Why is this?
4. What two substances are created when an acid is mixed with a base?
5. What are the tastes of acids and bases?
6. How does a base feel? How does that compare to the texture of acids?
7. What type of gas is released when an acid mixes with a metal (assuming that there is a reaction)?
8. Acids and bases break down into charge ions called _____ that can conduct electricity. Does strong or weak acid/base conduct more electricity?
9. What is an indicator? What are the 3 types of indicators in the form of paper? What are some indicators in the form of solution?
10. What indicator solution is clear in an acid, but turns pink in a base ($\text{pH} > 8$)?
11. What is the ion product (number) for water at 25 $^{\circ}\text{C}$?
12. What is an amphoteric substance?
13. What determines the strength of an acid? List strong acids and weak acids.
14. How is the strength of a base determined? List strong bases and weak bases.
15. Give the formulas of the following acids: hydrochloric acid, nitrous acid, acetic acid, carbonic acid, and sulfuric acid.
16. What are the names of the following acids? HNO_3 , HF , H_2PO_4 , H_2SO_3 .
17. What color does blue litmus paper turn in an acidic solution? A basic solution?
18. What is the range of pH for an acid or a base of the known concentration?
19. What is a solution with a pH=7 called? Give an example of a solution with this pH.
20. Describe the concept of conjugate acid-base pair. What is the different between acid and conjugate acid? Base and conjugate base? Show an example.
21. What is a buffer solution?
22. What type of solution contains an acid or base of a known concentration in titration?
23. What is the pH of a solution where $[\text{OH}^-]$ is higher than $[\text{H}_3\text{O}^+]$? Give a range.
24. If the $[\text{H}_3\text{O}^+]$ is greater than $[\text{OH}^-]$, is the solution acidic or basic?
25. What is the end point of titration? How is this different from equivalence point?
26. What should the end point of titration be close to?
27. What is a buffer capacity?
28. If $[\text{H}_3\text{O}^+]$ or $[\text{OH}^-]$ are added beyond the buffer capacity, what will happen to the pH?

Problems:

29. If pH is 4, what is pOH?
30. In a solution of water, the concentration of OH^- ions is $1.0 \times 10^{-6}\text{M}$. What is the concentration of the hydronium ions?
31. The concentration of OH^- ions in a solution is $3.4 \times 10^{-7}\text{ M}$. Is the solution acidic, basic or neutral? What is the pH of the solution?
32. A volume of 37 mL of 0.36 M KCN neutralizes a 75-mL sample of HClO solution (acid). What is the concentration of HClO ?

Acid And Bases Study Guide Ch 15

David Margolese,David W. Oxtoby

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The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills,Amina El-Ashmawy,2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone,Theodore O. Dolter,2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections assessment exercises at the end each section and relevant chapter problems at the end of each chapter A new Math Check allows quick access to the needed basic skill The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem Every concept in the text is clearly illustrated with one or more step by step examples Making it Real essays have been updated to present timely and engaging real world applications emphasizing the relevance of the material they are learning This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter

General Chemistry Study Guide Sixth Edition Darrell D. Ebbing,1999

Pediatric Critical Care Study Guide Steven E. Lucking,Frank A. Maffei,Robert F.

Tamburro,Neal J. Thomas,2012-04-10 This is the first comprehensive study guide covering all aspects of pediatric critical care medicine It fills a void that exists in learning resources currently available to pediatric critical care practitioners The major textbooks are excellent references but do not allow concise reading on specific topics and are not intended to act as both text and study guide There are also several handbooks available but these are usually written for general pediatric residents and lack the advanced physiology and pathophysiology required for the higher level pediatric critical care practitioner

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Kenneth L. Henold, 1984 **Student Study Guide to Accompany Petrucci's General Chemistry** Robert K. Wissner, 1985

Study Guide with Selected Solutions Karen Timberlake, 2002-08-06 Written by the author the Study Guide is keyed to the learning goals in the text and designed to promote active learning through a variety of exercises with answers and mastery exams Also contains complete solutions to odd numbered problems *Ess Chem Probs Study Guide* David Margolese, David W. Oxtoby, 2004-04 **Study Guide and Solutions Manual for Seager/Slabaugh's Chemistry for Today** Seager/Slabaugh, 2004 The fifth edition of the Study Guide and Student Solutions Manual has been updated to reflect all of the changes to the text This ancillary tests the student on the learning objectives in each chapter and provides answers to all of the even numbered end of chapter exercises New additional activities have been added to include a review of each section of the chapter and a section entitled Tying It All Together with a Laboratory Application **Student Study Guide to Accompany Chemistry** Kenneth J. Hughes, Paul B. Kelter, 2002-05 Reviews essential points from the text presents problems and hints to assist the student in solving them Solutions to selected problems in the end of chapter problem sets are included By Kenneth Hughes of the University of Wisconsin Madison *Study Guide to Organic Chemistry, 4th Ed*

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