

Acids and Bases

MERCK

Acid			Base			
Strength of Acid	Ka	Name	Formula	Formula	Name	Strength of Base
Strongest	Large	Perchloric acid	HClO ₄	ClO ₄ ⁻	Perchlorate ion	
	3.2×10^{-1}	Hydroiodic acid	HBr	I ⁻	Iodide	
	1.0×10^{-1}	Hydrobromic acid	HCl	Br ⁻	Bromide	
	1.3×10^{-1}	Hydrochloric acid	H ₂ SO ₄	Cl ⁻	Chloride	
	1.0×10^{-2}	Sulfuric acid	HNO ₃	HSO ₄ ⁻	Hydrogen sulfate ion	
	2.4×10^{-3}	Nitric acid	H ₃ O ⁺	NO ₃ ⁻	Nitrate ion	
		Hydroxium ion		H ₂ O	Water	
	5.4×10^{-4}	oxalic acid	HO ₂ C ₂ O ₄ H	HO ₂ C ₂ O ₄ ²⁻	Hydrogen oxalate ion	
	1.3×10^{-4}	Sulfurous acid	H ₂ SO ₃	HSO ₃ ⁻	Hydrogen sulfite ion	
	1.0×10^{-4}	Hydrogen sulfate ion	HSO ₄ ⁻	SO ₄ ²⁻	Sulfate ion	
	7.1×10^{-5}	Phosphoric acid	H ₃ PO ₄	H ₂ PO ₄ ⁻	Dihydrogen phosphate ion	
	7.2×10^{-5}	Nitrous acid	HNO ₂	NO ₂ ⁻	Nitrite ion	
	6.6×10^{-6}	Hydrofluoric acid	HF	F ⁻	Fluoride ion	
	1.8×10^{-6}	Methanolic acid	HCO ₂ H	HCO ₃ ⁻	Methanoate ion	
	6.3×10^{-7}	Benzoic acid	C ₆ H ₅ COOH	C ₆ H ₅ COO ⁻	Benzoate ion	
	5.4×10^{-7}	Hydrogen oxalate ion	HO ₂ C ₂ O ₄ ²⁻	O ₂ C ₂ O ₄ ²⁻	Oxalate ion	
	1.8×10^{-7}	Ethanoic acid	CH ₃ COOH	CH ₃ COO ⁻	Ethanoate (acetate) ion	
	4.4×10^{-7}	Carbonic acid	H ₂ CO ₃	HCO ₃ ⁻	Hydrogen carbonate ion	
	1.1×10^{-7}	Hydrosulfuric acid	H ₂ S	HS ⁻	Hydrogen sulfide ion	
	6.3×10^{-8}	Dihydrogen phosphate ion	H ₂ PO ₄ ⁻	HPO ₄ ²⁻	Hydrogen phosphate ion	
	6.2×10^{-8}	Hydrogen sulfite ion	HS ⁻	SO ₃ ²⁻	Sulfite ion	
	2.9×10^{-8}	Hypochlorous acid	HClO	ClO ⁻	Hypochlorite ion	
	6.2×10^{-10}	Hydrocyanic acid	HCN	CN ⁻	Cyanide ion	
	5.8×10^{-10}	Ammonium ion	NH ₄ ⁺	NH ₃	Ammonia	
	5.8×10^{-10}	Boric acid	H ₃ BO ₃	H ₂ BO ₃ ⁻	Dihydrogen borate	
	4.7×10^{-11}	Hydrogen carbonate ion	HCO ₃ ⁻	CO ₃ ²⁻	Carbonate ion	
	4.2×10^{-11}	Hydrogen phosphate ion	HPO ₄ ²⁻	PO ₄ ³⁻	Phosphate ion	
	1.8×10^{-12}	Dihydrogen borate ion	H ₂ BO ₃ ⁻	HBO ₃ ²⁻	Hydrogen borate ion	
	1.3×10^{-12}	Hydrogen sulfide ion	HS ⁻	S ²⁻	Sulfide ion	
	1.6×10^{-12}	Hydrogen borate ion	HBO ₃ ²⁻	BO ₃ ³⁻	Borate ion	
Weakest	-----	water	H ₂ O	OH ⁻	Hydroxide	Strongest



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Tina Youngman

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1. Understanding the eBook Acids And Bases Guide
 - The Rise of Digital Reading Acids And Bases Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Acids And Bases Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acids And Bases Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acids And Bases Guide
 - Personalized Recommendations
 - Acids And Bases Guide User Reviews and Ratings
 - Acids And Bases Guide and Bestseller Lists
5. Accessing Acids And Bases Guide Free and Paid eBooks

- Acids And Bases Guide Public Domain eBooks
- Acids And Bases Guide eBook Subscription Services
- Acids And Bases Guide Budget-Friendly Options

6. Navigating Acids And Bases Guide eBook Formats

- ePUB, PDF, MOBI, and More
- Acids And Bases Guide Compatibility with Devices
- Acids And Bases Guide Enhanced eBook Features

7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of Acids And Bases Guide
- Highlighting and Note-Taking Acids And Bases Guide
- Interactive Elements Acids And Bases Guide

8. Staying Engaged with Acids And Bases Guide

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Acids And Bases Guide

9. Balancing eBooks and Physical Books Acids And Bases Guide

- Benefits of a Digital Library
- Creating a Diverse Reading Collection Acids And Bases Guide

10. Overcoming Reading Challenges

- Dealing with Digital Eye Strain
- Minimizing Distractions
- Managing Screen Time

11. Cultivating a Reading Routine Acids And Bases Guide

- Setting Reading Goals Acids And Bases Guide
- Carving Out Dedicated Reading Time

12. Sourcing Reliable Information of Acids And Bases Guide

- Fact-Checking eBook Content of Acids And Bases Guide
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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