

Acids and Bases

MERCK

Acid				Base		
Strength of Acid	Ka	Name	Formula	Formula	Name	Strength of Base
Strongest 	Large	Perchloric acid	HClO_4	ClO_4^-	Perchlorate ion	Weakest 
	3.2×10^8	Hydroiodic acid		I^-	Iodide	
	1.0×10^9	Hydrobromic acid	HBr	Br^-	Bromide	
	1.3×10^4	Hydrochloric acid	HCl	Cl^-	Chloride	
	1.0×10^3	Sulfuric acid	H_2SO_4	HSO_4^-	Hydrogen sulfate ion	
	2.4×10^1	Nitric acid	HNO_3	NO_3^-	Nitrate ion	
	-----	Hydronium ion	H_3O^+	H_2O	Water	
	5.4×10^{-2}	Oxalic acid	$\text{HO}_2\text{C}_2\text{O}_2\text{H}$	$\text{HO}_2\text{C}_2\text{O}_2^-$	Hydrogen oxalate ion	
	1.3×10^{-2}	Sulfurous acid	H_2SO_3	HSO_3^-	Hydrogen sulfite ion	
	1.0×10^{-2}	Hydrogen sulfate ion	HSO_4^-	SO_4^{2-}	Sulfate ion	
Medium 	7.1×10^{-3}	Phosphoric acid	H_3PO_4	H_2PO_4^-	Dihydrogen phosphate ion	Medium 
	7.2×10^{-4}	Nitrous acid	HNO_2	NO_2^-	Nitrite ion	
	6.6×10^{-4}	Hydrofluoric acid	HF	F^-	Fluoride ion	
	1.8×10^{-4}	Methanoic acid	HCO_2H	HCO_2^-	Methanoate ion	
	6.3×10^{-5}	Benzoic acid	$\text{C}_6\text{H}_5\text{COOH}$	$\text{C}_6\text{H}_5\text{COO}^-$	Benzoate ion	
	5.4×10^{-5}	Hydrogen oxalate ion	$\text{HO}_2\text{C}_2\text{O}_2^-$	$\text{O}_2\text{C}_2\text{O}_2^{2-}$	Oxalate ion	
	1.8×10^{-5}	Ethanoic acid	CH_3COOH	CH_3COO^-	Ethanoate (acetate) ion	
	4.4×10^{-7}	Carbonic acid	H_2CO_3	HCO_3^-	Hydrogen carbonate ion	
	1.1×10^{-7}	Hydrosulfuric acid	H_2S	HS^-	Hydrogen sulfide ion	
	6.3×10^{-8}	Dihydrogen phosphate ion	H_2PO_4^-	HPO_4^{2-}	Hydrogen phosphate ion	
Weakest	6.2×10^{-8}	Hydrogen sulfite ion	HS^-	SO_3^{2-}	Sulfite ion	Strongest
	2.9×10^{-8}	Hypochlorous acid	HClO	ClO^-	Hypochlorite ion	
	6.2×10^{-10}	Hydrocyanic acid	HCN	CN^-	Cyanide ion	
	5.8×10^{-10}	Ammonium ion	NH_4^+	NH_3	Ammonia	
	5.8×10^{-10}	Boric acid	H_3BO_3	H_2BO_3^-	Dihydrogen borate	
	4.7×10^{-11}	Hydrogen carbonate ion	HCO_3^-	CO_3^{2-}	Carbonate ion	
	4.2×10^{-13}	Hydrogen phosphate ion	HPO_4^{2-}	PO_4^{3-}	Phosphate ion	
	1.8×10^{-13}	Dihydrogen borate ion	H_2BO_3^-	HBO_3^{2-}	Hydrogen borate ion	
	1.3×10^{-13}	Hydrogen sulfide ion	HS^-	S^{2-}	Sulfide ion	
	1.6×10^{-14}	Hydrogen borate ion	HBO_3^{2-}	BO_3^{3-}	Borate ion	
	-----	water	H_2O	OH^-	Hydroxide	



Acids And Bases Guide

Tina Youngman



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Guide to Acids & Bases Kit Centre for Adult Education, Shell South Africa (Pty) Limited, 1994 **Pocket Guide to Fluid, Electrolyte, and Acid-Base Balance - E-Book** Ursula Heitz, Mima M. Horne, 2004-10-20 Compares different types of parenteral and enteral feedings along with nursing implications Contains expanded and updated relevant IV therapy and pharmacology Features updated content throughout with current literature and research findings such as most current blood pressure guidelines from the U S Department of Health and Human Services **Handbook of Acid-Base Indicators** R. W. Sabnis, 2007-10-04 While acid base indicators continue to find new applications in an ever widening range of scientific disciplines there is no current book that focuses entirely on the subject nor one that brings together the relevant advances that have evolved over the last three decades The Handbook of Acid Base Indicators compiles the most up to date c
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Table of Contents Acids And Bases Guide

1. Understanding the eBook Acids And Bases Guide
 - The Rise of Digital Reading Acids And Bases Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Acids And Bases Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Acids And Bases Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Acids And Bases Guide
 - Personalized Recommendations
 - Acids And Bases Guide User Reviews and Ratings
 - Acids And Bases Guide and Bestseller Lists
5. Accessing Acids And Bases Guide Free and Paid eBooks

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- Acids And Bases Guide eBook Subscription Services
- Acids And Bases Guide Budget-Friendly Options
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 - ePub, PDF, MOBI, and More
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 - Acids And Bases Guide Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Acids And Bases Guide
 - Highlighting and Note-Taking Acids And Bases Guide
 - Interactive Elements Acids And Bases Guide
- 8. Staying Engaged with Acids And Bases Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Acids And Bases Guide
- 9. Balancing eBooks and Physical Books Acids And Bases Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Acids And Bases Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Acids And Bases Guide
 - Setting Reading Goals Acids And Bases Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Acids And Bases Guide
 - Fact-Checking eBook Content of Acids And Bases Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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