

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a given acid or hydrogen ion concentration.
2. Calculate hydrogen ion concentration from pH and hydroxide ion concentration from pOH.
3. Calculate the pH of a weak acid.

Vocabulary

- | | | |
|-------------------|-------------------------------|------------------|
| 1. autoionization | 2. hydrogen ion concentration | 3. acid solution |
| 4. autoionization | 5. pOH | 6. acid solution |

Key Equation

- | | |
|-------------------------------------|-------------------------------------|
| 1. $\text{pH} = -\log [\text{H}^+]$ | 2. $\text{pH} = -\log [\text{H}^+]$ |
| 3. $\text{pH} = -\log [\text{H}^+]$ | 4. $\text{pH} = -\log [\text{H}^+]$ |

Part A Completion

Write the hydrogen ion concentration or autoionization constant, if the missing quantity that is indicated in the partial equation below represents either an autoionization constant.

Water autoionizes to H^+ and hydroxide ions. The autoionization constant K_w of the autoionization of water which produces an H^+ ion and a hydroxide ion is 1.0×10^{-14} .

Hydrochloric acid dissociates from HCl to H^+ and Cl^- . The K_a of HCl is 1.0×10^{-7} . The concentration of H^+ in the solution is 1.0×10^{-7} . The concentration of Cl^- is 1.0×10^{-7} .

HCl is a strong acid. The K_a of HCl is 1.0×10^{-7} . The concentration of H^+ in the solution is 1.0×10^{-7} . The concentration of Cl^- is 1.0×10^{-7} .

- | |
|-----------|
| 1. K_w |
| 2. K_a |
| 3. K_b |
| 4. K_w |
| 5. K_a |
| 6. K_b |
| 7. K_w |
| 8. K_a |
| 9. K_b |
| 10. K_w |
| 11. K_a |
| 12. K_b |

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Huangqi Zhang



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