

## ACIDS, BASES, & THE pH SCALE

Name: **Answer Key**

Date: \_\_\_\_\_

Pd: \_\_\_\_\_

### PART 1: ACIDS AND BASES

Comparing Acids and Bases	
Acids	Bases (Alkalines)
<ul style="list-style-type: none"><li>• <b>Sour</b> taste, leaves <b>burning</b> sensation</li><li>• <b>Increase</b> hydrogen ions (<math>H^+</math>) in solution</li><li>• Example: <b>HCl</b>     <b><math>H^+ + Cl^-</math></b></li><li>• Turns <b>litmus</b> (pH) paper to <b>red</b></li></ul>	<ul style="list-style-type: none"><li>• Sharp <b>bitter</b> taste, feels <b>slippery</b> to touch</li><li>• <b>Decrease</b> hydrogen ions (<math>H^+</math>) in solution (<b>Increase</b> hydroxide ions (<math>OH^-</math>) in solution)</li><li>• Example: <b>NaOH</b>     <b><math>Na^+ + OH^-</math></b></li><li>• Turns litmus (pH) paper to <b>blue</b></li></ul>

### PART 2: THE pH SCALE

- The **pH Scale** is a **measurement** of how acidic or how basic a solution is.
- The scale goes from **0** to **14**.
- **Compounds** are **acidic** if they have a pH **lower** than **7**.
- Compounds with a pH **higher** than **7** are **basic** or **alkaline**.
- Compounds with a pH **of 7**, are **neutral**.

0                      1           2           3           4           5           6           7           8           9           10           11           12           13                      14

**Increased** concentration of  **$H^+$**  ions  
**Decreased** concentration of  **$OH^-$**  ions

**Decreased** concentration of  **$H^+$**  ions  
**Increased** concentration of  **$OH^-$**  ions

pH below 7 = **acids**

pH at 7 = **neutral**  
(**pure water**)

pH above 7 = **bases**

### PART 3: BUFFERS

- A **buffer** is a **weak acid** or **base** that **prevents** sudden **changes** in pH.
  - Maintains **homeostasis**
- A buffer works by **binding** to  $H^+$  ions **releasing**  $H^+$  ions when the solution pH changes.
- Buffers **bind** to  $H^+$  ions when the **concentration** of  $H^+$  ions in solution **increases**.
- **NaOH**      **$Na^+ + OH^-$**  (The  $OH^-$  will bind to  $H^+$  ions to form  $H_2O$  – **neutralization**)
- Buffers **release**  $H^+$  ions when the **concentration** of  $H^+$  ions in solution **decreases**.  
**HCl**      **$H^+ + Cl^-$**

# Acids And Bases Answer Key

**William R. Robinson, Jerome D.  
Odom, Henry Fuller Holtzclaw**



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