

ACIDS, BASES, & THE pH SCALE

Name: *Answer Key*

Date:

Pd:

PART 1: ACIDS AND BASES

Comparing Acids and Bases	
Acids	Bases (Alkalines)
• Sour taste, leaves burning sensation	• Sharp bitter taste, feels slippery to touch
• Increase hydrogen ions (H^+) in solution	• Decrease hydrogen ions (H^+) in solution (Increase hydroxide ions (OH^-) in solution)
• Example: HCl $H^+ + Cl^-$	• Example: $NaOH$ $Na^+ + OH^-$
• Turns litmus (pH) paper to red	• Turns litmus (pH) paper to blue

PART 2: THE pH SCALE

- The pH Scale is a measurement of how acidic or how basic a solution is.
- The scale goes from 0 to 14.
- Compounds are acidic if they have a pH lower than 7.
- Compounds with a pH higher than 7 are basic or alkaline.
- Compounds with a pH of 7 are neutral.



Increased concentration of H^+ ions

Decreased concentration of OH^- ions

pH below 7 = acids

pH at 7 = neutral
(pure water)

Decreased concentration of H^+ ions

Increased concentration of OH^- ions

pH above 7 = bases

PART 3: BUFFERS

- A buffer is a weak acid or base that prevents sudden changes in pH.
 - Maintains homeostasis
- A buffer works by binding to H^+ ions releasing H^+ ions when the solution pH changes.
- Buffers bind to H^+ ions when the concentration of H^+ ions in solution increases.
- $NaOH$ $Na^+ + OH^-$ (The OH^- will bind to H^+ ions to form H_2O – neutralization)
- Buffers release H^+ ions when the concentration of H^+ ions in solution decreases.
 HCl $H^+ + Cl^-$

Acids And Bases Answer Key

**William R. Robinson, Jerome D.
Odom, Henry Fuller Holtzclaw**

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