

14.3

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Calculate concentrations based on the pH of a solution and the pH of a solution based on concentration.
2. Calculate pH and concentration of acids and bases and the pH of a solution based on concentration.
3. Calculate the pH of a solution (25°C).

Vocabulary

- | | | |
|-------------------|---|------------------|
| 1. autoionization | 2. hydronium ion (H_3O^+) | 3. acid solution |
| 4. autoionization | 5. acid rain | 6. acid rain |
| 7. autoionization | 8. acid rain | 9. acid |

Key Equation

- | | |
|------------------------------------|--|
| 1. $\text{pH} = -\log[\text{H}^+]$ | 2. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |
| 3. $\text{pH} = -\log[\text{H}^+]$ | 4. $\text{pH} = -\log[\text{H}_3\text{O}^+]$ |

Part A Completion

Write the equation for the autoionization of water. What is the concentration of H^+ in pure water? What is the concentration of OH^- in pure water?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ? What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ OH^- ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ? What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ OH^- ? What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ?

What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ? What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ OH^- ? What is the pH of a solution with a concentration of $1.0 \times 10^{-7} \text{ M}$ H^+ ?

1. $\text{H}_2\text{O} \rightleftharpoons \text{H}^+ + \text{OH}^-$ 2. $1.0 \times 10^{-7} \text{ M}$ 3. $1.0 \times 10^{-7} \text{ M}$ 4. $1.0 \times 10^{-7} \text{ M}$ 5. $1.0 \times 10^{-7} \text{ M}$ 6. $1.0 \times 10^{-7} \text{ M}$ 7. $1.0 \times 10^{-7} \text{ M}$ 8. $1.0 \times 10^{-7} \text{ M}$ 9. $1.0 \times 10^{-7} \text{ M}$ 10. $1.0 \times 10^{-7} \text{ M}$ 11. $1.0 \times 10^{-7} \text{ M}$

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Table of Contents 19 2 Hydrogen Ions And Acidity Answer Key

1. Understanding the eBook 19 2 Hydrogen Ions And Acidity Answer Key
 - The Rise of Digital Reading 19 2 Hydrogen Ions And Acidity Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying 19 2 Hydrogen Ions And Acidity Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 19 2 Hydrogen Ions And Acidity Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from 19 2 Hydrogen Ions And Acidity Answer Key
 - Personalized Recommendations
 - 19 2 Hydrogen Ions And Acidity Answer Key User Reviews and Ratings
 - 19 2 Hydrogen Ions And Acidity Answer Key and Bestseller Lists
5. Accessing 19 2 Hydrogen Ions And Acidity Answer Key Free and Paid eBooks
 - 19 2 Hydrogen Ions And Acidity Answer Key Public Domain eBooks
 - 19 2 Hydrogen Ions And Acidity Answer Key eBook Subscription Services
 - 19 2 Hydrogen Ions And Acidity Answer Key Budget-Friendly Options
6. Navigating 19 2 Hydrogen Ions And Acidity Answer Key eBook Formats

- ePub, PDF, MOBI, and More
 - 19 2 Hydrogen Ions And Acidity Answer Key Compatibility with Devices
 - 19 2 Hydrogen Ions And Acidity Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of 19 2 Hydrogen Ions And Acidity Answer Key
 - Highlighting and Note-Taking 19 2 Hydrogen Ions And Acidity Answer Key
 - Interactive Elements 19 2 Hydrogen Ions And Acidity Answer Key
 8. Staying Engaged with 19 2 Hydrogen Ions And Acidity Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 19 2 Hydrogen Ions And Acidity Answer Key
 9. Balancing eBooks and Physical Books 19 2 Hydrogen Ions And Acidity Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 19 2 Hydrogen Ions And Acidity Answer Key
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 19 2 Hydrogen Ions And Acidity Answer Key
 - Setting Reading Goals 19 2 Hydrogen Ions And Acidity Answer Key
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 19 2 Hydrogen Ions And Acidity Answer Key
 - Fact-Checking eBook Content of 19 2 Hydrogen Ions And Acidity Answer Key
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
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