

In the figure, a block of mass $m = 12 \text{ kg}$ is released from rest on a frictionless incline of angle $\theta = 30^\circ$. Below the block is a spring that can be compressed 2.0 cm by a force of 270 N. The block momentarily stops when it compresses the spring by 5.5 cm. (a) How far does the block move down the incline from its rest position to this stopping point? (b) What is the speed of the block just as it touches the spring?

$$\cancel{K_A} + U_A = \cancel{K_C} + U_C$$

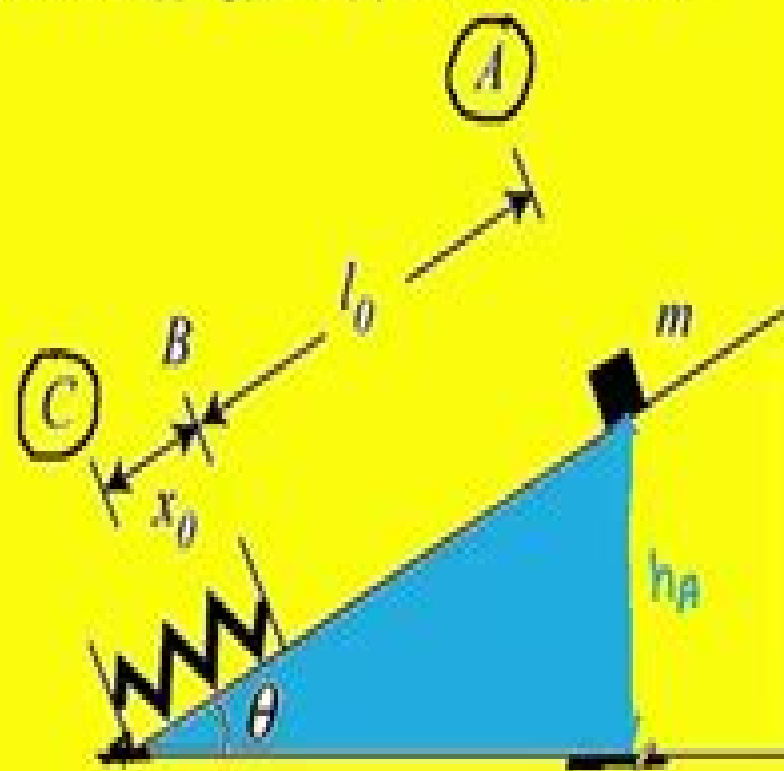
$$h_A = \frac{\frac{1}{2} K (\Delta x)^2}{mg}$$

$$\sin \theta = \frac{h_A}{(l_0 + x_0)}$$

$$= \frac{\frac{1}{2} (13500) (0.055)^2}{12 (4.8)}$$

$$13500 \text{ N/m} = K$$

$$h_A = 0.174 \text{ m}$$



Answers Pogil How Does A Battery Work

Adele Richardson



Answers Pogil How Does A Battery Work:

How Batteries Work Victoria G. Christensen, 2016-08-01 Have you ever wondered how a set of batteries power the remote for your TV Or have you ever wanted to examine the battery inside a car Without batteries we would not be able to power many of the electronics we use every day But how does a battery store energy and provide electricity to these devices What are the parts that make a battery work Discover the history of batteries and how scientists and engineers figured out ways to store electricity Explore what makes up a battery including an anode a cathode and an electrolyte and learn how batteries power all kinds of electronics from a cell phone to a bus

How Does a Battery Work? Rosie Wilson, 2013-01-01 Batteries are amazing little bundles of energy We rarely think about how they work when we use them in devices This book explains how batteries convert chemical energy into electrical energy and how they power so many things in our world Readers also learn through this accessible text about different types of batteries and the importance of recycling these energy sources

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The Book of Battery Manufacturing Razzab Ali Shaikh, 2014-10-06 What are the main parts of a battery All batteries contain one or more cells but people often use the terms battery and cell interchangeably A cell is just the working chemical unit inside a battery one battery can contain any number of cells A cell has three main parts a positive electrode terminal a negative electrode and a liquid or solid separating them called the electrolyte When a battery is connected to an electric circuit a chemical reaction takes place in the electrolyte causing ions in this case atoms with a positive electrical charge to flow through it one way with electrons particles with a negative charge flowing through the outer circuit in the other direction This movement of electric charge makes an electric current flow through the cell and through the circuit it is connected to That s the theory anyway Now let s look at it in practice How does a battery really work Where does the power in a battery actually come from Let s take a closer look Here s my battery hooked up to a flashlight bulb to make a simple circuit I ve unwrapped a paperclip to make a piece of connecting wire and I m holding that between the bottom of the battery and the side of the bulb If you look closely you can see the bulb is shining That s because electrons are marching through it Anode and cathode Now here s what s going on inside The battery s positive terminal shown just above my left thumb in the photo and colored red in the artwork below is connected to a positive electrode that s mostly hidden inside the battery We call this the cathode The outer case and the bottom of the battery make up the negative terminal or negative electrode which is also called the anode and colored green in the artwork The paperclip wire is represented in the art by the blue line Let s quickly clear up one point of confusion At school you may have learned that the cathode is the negative electrode and the anode the positive electrode However that

applies only to electrolysis passing electricity through a chemical to split it up Batteries are like electrolysis going backwards they split up chemicals to make electricity so the terms anode and cathode are switched around Okay To avoid confusion I suggest it's best not to use the terms anode and cathode at all It's better to say positive terminal and negative terminal and then it's always clear what you mean whether you're talking about batteries or electrolysis Chemical reactions Now back to our battery The positive and negative electrodes are separated by the chemical electrolyte It can be a liquid but in an ordinary battery it is more likely to be a dry powder When you connect the battery to a lamp and switch on chemical reactions start happening One of the reactions generates positive ions shown here as big yellow blobs and electrons smaller brown blobs at the negative electrode The positive ions flow through the electrolyte to the positive electrode from the green line to the red one Meanwhile the electrons smaller brown blobs flow around the outside circuit blue line to the positive electrode and make the lamp light up on the way The electrons and ions flow because of the chemical reactions happening inside the battery usually two or three of them going on simultaneously The exact reactions depend on the materials from which the electrodes and electrolyte are made and we won't go into them here If you want to know what they are Google the type of the battery you're interested in followed by the words anode cathode reactions For example Google zinc carbon battery anode cathode reactions Whatever chemical reactions take place the general principle of electrons going around the outer circuit and ions flowing in the opposite direction through the electrolyte happens in all batteries As the battery generates power the chemicals inside it are gradually converted into different chemicals

Electricity Adele Richardson, 2006 Introduces electricity and its generation components movement and functions in a question and answer format Provided by publisher

Storage Batteries A. E. Watson, 2017-06-22 Excerpt from Storage Batteries Their Theory Construction and Use A storage battery does not store electricity It is a device for storing chemical energy by the action of an electric current Of course the additional qualifications are inferred that the storage can be maintained for an indefinite time and at will under favorable conditions this chemical change may be allowed to undo itself and return as it were the very current that produced it A crude illustration of a storage of energy may be seen in a stick of wood Energy has been expended in producing it and until used the energy is in the potential form that is it is ready and willing to do work When the stick is used for fuel the energy becomes manifest in the form of heat and in the language of the physicist it is now kinetic If now from the ashes smoke and gases and with an expenditure Of energy just equivalent to the heat given out in burning the stick be reconstructed an artificial storage of energy would have been made Except for the very slow burning called rot or decay the storage could be again conceived as Of indefinite duration About the Publisher Forgotten Books publishes hundreds of thousands of rare and classic books Find more at www.forgottenbooks.com This book is a reproduction of an important historical work Forgotten Books uses state of the art technology to digitally reconstruct the work preserving the original format whilst repairing imperfections present in the aged copy In rare cases an imperfection in the original such as a blemish

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