



that consists of the basic, the components that are present in all matter and cannot be split. In simple terms, that form part of all things around us.

What is an element?

At 0°C , 100 g of H_2O

What is an atom?

If you look at an element, you cannot see it. You can only see the atoms that make it up. After all, an element is just one type of atom.

At 0°C , 100 g of H_2O



At 100°C , 100 g of H_2O consists of water molecules. These molecules are moving faster than they were at 0°C . The water molecules are moving faster because the temperature is higher. This is because the temperature is a measure of the average kinetic energy of the particles in a substance.

- It contains two H atoms and a O atom, what is it called?
- Water is formed in the chemical equation $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$. Which side of the equation has the reactants, the left side or the right side?
- Which side of the equation has the products, the left side or the right side?
- What is the most stable electron configuration for an atom?
 - Energy level has an equal number of electrons.
 - The highest energy level has a full set of electrons.
 - There is an even number of electrons.
- Which one of the following sets has an even number of valence electrons?
 - Sharing electrons with another atom.
 - Transferring electrons to another atom.
 - Adding energy to an existing atom.
- Which type of bond forces atoms to move closer together?
 - Ionic bond.
 - Covalent bond.
 - Hydrogen bond.
- When two atoms have an unequal pull on the electrons, they share the bond. This is called a polar bond.
- Parts of the following bond types are the bonds?
 - Ionic bonds.
 - Covalent bonds.
 - Hydrogen bonds.
- Which one of these is shown using H_2O ?

- Which property of water is not a result of hydrogen bonding?
 - cohesive nature
 - transparency
 - viscosity
 - adhesion

- A water molecule is able to split into two oxygen atoms because water has **homopolar bonds**. This results in a phenomenon called **depolarization** because it breaks the bond and the two oxygen atoms move away from each other.

ADVANCED CONCEPTS

- In the following chemical equation, how many molecules of water are formed in a reaction?

$$2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$$
- In the chemical equation in question 13, how many molecules of oxygen are present in a reaction?
- Write how many atoms of each element are present in each side of the following equation and determine which element is present in greatest and the smallest amount.

$$\text{NaCl} \rightarrow \text{NaCl} + \text{O}_2$$
- How would you balance the reaction in question 13?

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Samuel Austin Allibone

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