

**CHAPTER 15** STUDY GUIDE FOR CONTENT MASTERY**Solutions****Section 15.1 What are solutions?**

In your textbook, read about the characteristic of solutions.

Underline the terms below just now to complete the passage.

solvent	liquid	solvent	solvent
insoluble	solid	solid	solid

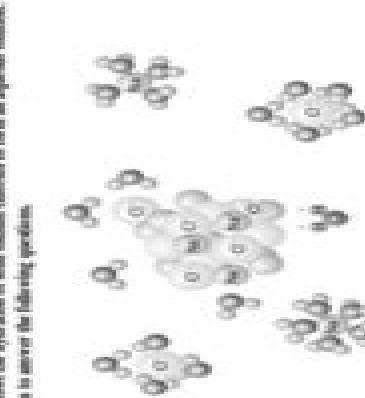
After reading (1) \_\_\_\_\_ solution \_\_\_\_\_ of solute is dissolved in solvent.

(2) \_\_\_\_\_ solute \_\_\_\_\_ is the (3) \_\_\_\_\_ solvent \_\_\_\_\_ and solute is the

(4) \_\_\_\_\_ solution. A solution that does not dissolve is  
 (5) \_\_\_\_\_ insoluble. (6) \_\_\_\_\_ liquid \_\_\_\_\_ solution is the most common  
 type of solution. If one liquid is soluble in another liquid, such as water and oil, the  
 two liquids are (7) \_\_\_\_\_ miscible. However, if one liquid is insoluble in another  
 the liquids are (8) \_\_\_\_\_ immiscible.

Read about solubility in aqueous solution in your textbook.

Underline the questions to answer the following questions.



The diagram shows the hydration of solid sodium chloride to form an aqueous solution.

Underline the diagram to answer the following questions.

1. In a solution, what is the solute? What is the solvent?  
 2. In a solution, what is the process of surrounding solute particles with solvent particles to form a solution?

**CHAPTER 15 STUDY GUIDE FOR CONTENT MASTERY****Section 15.1** Section**STUDY GUIDE FOR CONTENT MASTERY**

16. An sodium chloride dissolved in water. What happens to the sodium and chloride ions?  
 The sodium and chloride ions are **separated and surrounded by the water molecules.**

17. Explain the attraction of the water molecules around the sodium ions and chloride ions.  
 Because the sodium ion is **positively charged**, it **attracts the negatively charged portion of the water molecule (the oxygen atom)** and **repels the positively charged portion of the water molecule (the hydrogen atom)**, because the chloride ion is **negatively charged**, it **attracts the positively charged portion of the water molecule** and **repels the negatively charged portion.**

18. How does the strength of the attraction between water molecules and sodium and chloride ions compare with the strength of the attraction between the sodium ions and chloride ions? How do you know?  
 The attraction between the **water molecules** and the **sodium and chloride ions** is **greater** than the attraction between the **sodium and chloride ions**. The **greater** strength of attraction between the **water molecules** and the **ions** is what **causes** the **solution process** to **occur**.

19. In a glass cup, add the rest of solution, stirring the solution, breaking the solution into smaller pieces, and heating the solution.

In your textbook, read about heat of solution, solubility and factors that affect solubility for each solute/solvent, write true or false.

True 20. The overall energy change that occurs when a solute dissolves is called the heat of solution.

False 21. Solubility is a measure of the maximum amount of solute the solution is a given amount of solvent at a specified temperature and pressure.

False 22. Solvent molecules in a glass of solution are in motion like the water molecules in a glass of water.

True 23. In a saturated solution, solute and crystallization are in equilibrium.

True 24. Additional solute can be dissolved in an unsaturated solution.

True 25. The solubility of a gas dissolved in a liquid decreases as the temperature of the solution increases.

**Section 15.2 Water and Oceans • Chapter 11****Study Guide for Content Mastery****Study Guide for Content Mastery**

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**Steven S. Zumdahl, Susan A. Zumdahl**

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