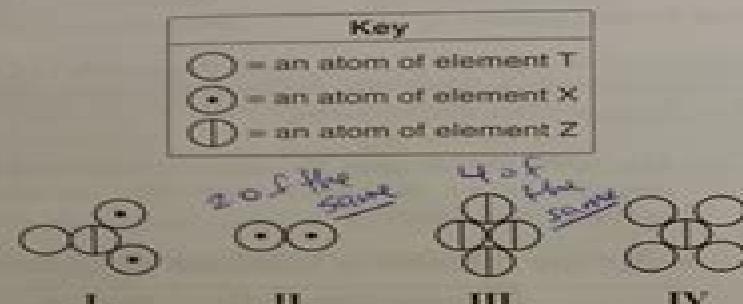


39 Given four particle models:



Which two models can be classified as elements?

(1) I and II      (3) II and III  
(2) I and IV      (4) II and IV

40 After being thoroughly stirred at 10°C, which mixture is heterogeneous?

(1) 25.0 g of KCl and 100. g of H<sub>2</sub>O ~~unsat.~~  
(2) 25.0 g of KNO<sub>3</sub> and 100. g of H<sub>2</sub>O ~~supersat.~~  
(3) 25.0 g of NaCl and 100. g of H<sub>2</sub>O ~~unsat.~~  
(4) 25.0 g of NaNO<sub>3</sub> and 100. g of H<sub>2</sub>O ~~unsat.~~

41 Which two compounds are electrolytes?

(1) KOH and CH<sub>3</sub>COOH  
(2) KOH and C<sub>2</sub>H<sub>12</sub>  
(3) CH<sub>3</sub>OH and CH<sub>3</sub>COOH  
(4) CH<sub>3</sub>OH and C<sub>2</sub>H<sub>12</sub>

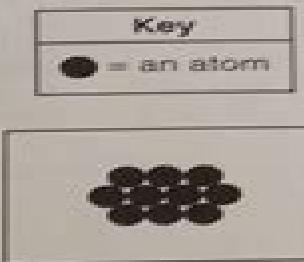
42 Which statement explains why a CO<sub>2</sub> molecule is nonpolar?

(1) Carbon and oxygen are both nonmetals.  
(2) Carbon and oxygen have different electronegativities.  
(3) The molecule has a symmetrical distribution of charge.  
(4) The molecule has an asymmetrical distribution of charge.

43 Which temperature change indicates an increase in the average kinetic energy of the molecules in a sample?

(1) 15°C to 298 K ↑      (3) 305 K to 0°C ↓  
(2) 37°C to 273 K ↓      (4) 355 K to 25°C ↓

44 Given the particle diagram:



Which substance at STP can be represented by this particle diagram? Solid

(1) N<sub>2</sub> (g)      (3) Mg(s)  
(2) H<sub>2</sub> (g)      (4) Kr (g)

45 Which type of equilibrium exists in a sealed flask containing Br<sub>2</sub>(l) and Br<sub>2</sub>(g) at 298 K and 1.0 atm?

(1) static phase equilibrium  
(2) static solution equilibrium  
(3) dynamic phase equilibrium  
(4) dynamic solution equilibrium

46 What are the products when potassium hydroxide reacts with hydrochloric acid?

Wash  
285  
(1) KH(s), Cl<sup>+</sup>(aq), and OH<sup>-</sup>(aq)  
(2) K(s), Cl<sub>2</sub>(g), and H<sub>2</sub>O(l)  
(3) KCl(aq) and H<sub>2</sub>O(l)  
(4) KOH(aq) and Cl<sub>2</sub>(g)

47 In a titration, 20.0 milliliters of a 0.150 M NaOH(aq) solution exactly neutralizes 24.0 milliliters of an HCl(aq) solution. What is the concentration of the HCl(aq) solution?

(1) 0.125 M      (3) 0.250 M  
(2) 0.180 M      (4) 0.360 M

48 What fraction of a Sr-90 sample remains unchanged after 87.3 years?

(1)  $\frac{1}{2}$       (3)  $\frac{1}{4}$   
(2)  $\frac{1}{3}$       (4)  $\frac{1}{8}$

3       $\frac{1}{2} = \frac{1}{4}$

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**Albert S. Tarendash**

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