

**Part – 1****Chapter 1 – Basic Concepts**

- 1) To find the number of molecules in a given compound which of the following is employed?  
 a) Chromatography  
 c) Kc  
 (NET-2015)
 

b) Le-Chatelier's principle  
 d) None of these
- 2) What is the percentage of N in  $\text{NH}_4\text{NO}_3$ ?  
 a) 40%  
 (NET-2015)
 

b) 50%  
 c) 35%  
 d) 30%
- 3) Mass spectrometer can be used to find:  
 a) Mass of atom  
 c) Charge to mass ratio of atom  
 (NET-2013)
 

b) Mass of ion  
 d) Charge to mass ratio of ion
- 4) 4 gram atoms of carbon has mass of:  
 a) 12  
 (NET-2013)
 

b) 24  
 c) 48  
 d) 56
- 5) 0.5 mole of  $\text{SO}_2$  has how much mass of oxygen?  
 a) 32  
 (NET-2013)
 

b) 12  
 c) 16  
 d) None of these
- 6) In combustion analysis method  $\text{CO}_2$  is absorbed by?  
 a)  $\text{Mg(OH)}_2$   
 (NET-2013)
 

b)  $\text{KOH}$   
 c)  $\text{Mg(ClO}_4)_2$   
 d) None of these
- 7) 1 amu is?  
 a)  $1.66 \times 10^{-27}$  kg  
 (NET-2013)
 

b)  $1.602 \times 10^{-19}$  kg  
 c)  $1.66 \times 10^{-27}$  kg  
 d) None of these
- 8) Mass of neutron in amu is?  
 a) 1.008386  
 (NET-2013)
 

b) 1.007376  
 c) 0.0005  
 d) None of these
- 9) Mass spectrometer is used to determine which of the following quantity?  
 a) Atomic masses  
 (NET-2013)
 

b) Atomic number  
 c) Both a and b  
 d) None of these
- 10) Chemical properties of an atom depends upon number of  
 a) Electrons  
 (NET-2013)
 

b) Neutrons  
 c) Protons  
 d) Nucleons
- 11) One mole of Oxygen has \_\_\_\_\_ number of atoms as compared to one mole of Nitrogen.  
 a) Equal  
 (NET-2013)
 

b) Half  
 c) Double  
 d) Quarter

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**Angie Papple Johnston, Rod Powers**



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